

CHAPMAN & HALL/CRC
CRYPTOGRAPHY AND NETWORK SECURITY

INTRODUCTION TO MODERN CRYPTOGRAPHY

Second Edition

Jonathan Katz
Yehuda Lindell



CRC Press
Taylor & Francis Group

A CHAPMAN & HALL BOOK

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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$1\text{K}^+ + 4\text{Al} + \text{NO}_3^- + 4\text{Al} + 3\text{NO}_3^- + 3\text{H}_2\text{O} + 3\text{H}^+$ ionic $3\text{Cl}^- + \text{Ag}^+ + 3\text{NO}_3^- + 4\text{Al} + 3\text{H}_2\text{O} + 3\text{H}^+$ net ionic $13\text{Cl}^- + \text{Ag}^+ + 4\text{Al} + 3\text{H}_2\text{O} + 3\text{H}^+$

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{ 2Ag}^+ \text{ 2F}^- \text{ 2NH}_4^+ \text{ Cr}_2\text{O}_7^{2-} \text{ Ag}_2\text{Cr}_2\text{O}_7 \text{ s } 2\text{NH}_4^+ \text{ 2F}^-$ aq because two $\text{NH}_4^+ \text{ aq}$

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