

# KINETICS Practice Problems and Solutions

Name: \_\_\_\_\_  
 Period: \_\_\_\_\_ Class: \_\_\_\_\_

AP Chemistry  
 Dr. McBride

The following questions represent general types of rate questions. Please answer each question completely and thoroughly. The solutions will be posted on the my Moodle page.

1. Consider the reaction:  $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$ . A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[\text{H}_2\text{O}]$ , mol/L	$[\text{O}_2]$ , mol/L	Initial Rate, mol/L·s
0.0100	0.0000	$1.20 \times 10^{-2}$
0.0200	0.0000	$4.80 \times 10^{-2}$
0.0300	0.0000	$8.10 \times 10^{-2}$

- a. Write the rate law for the reaction.  $\text{Rate} = k[\text{H}_2\text{O}]^2$   
 $k = 1.20 \times 10^{-2} \text{ s}^{-1}$
- b. Write the rate law for the reaction.  $\text{Rate} = k[\text{H}_2\text{O}]^2$
- c. Determine the order and units of the rate constant.  $k = 1.20 \times 10^{-2} \text{ s}^{-1}$  (order and units using the method in your textbook is not being tested on this question.)

2. Consider the reaction:  $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$ . A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[\text{H}_2\text{O}]$ , mol/L	$[\text{O}_2]$ , mol/L	Initial Rate, mol/L·s
0.010	0.000	0.010
0.020	0.000	0.040
0.030	0.000	0.090

- a. Write the rate law for the reaction.  $\text{Rate} = k[\text{H}_2\text{O}]^2$   
 $k = 1.00 \text{ s}^{-1}$
- b. Write the rate law for the reaction.  $\text{Rate} = k[\text{H}_2\text{O}]^2$
- c. Determine the order and units of the rate constant.  $k = 1.00 \text{ s}^{-1}$  (order and units using the method in your textbook is not being tested on this question.)

3. Consider the following reactions:  $\text{A}_2 + \text{B}_2 \rightarrow \text{A} + \text{B}$  (slow)  
 $\text{A}_2 + \text{B} \rightarrow \text{A} + \text{B}_2$  (fast)

- a. Write the overall balanced chemical equation:  $2\text{A}_2 + \text{B}_2 \rightarrow 2\text{A} + \text{B}_2$
- b. Identify any intermediates within the mechanism: B
- c. What is the order with respect to each reactant?  $\text{A}_2: 2; \text{B}_2: 0$

# Kinetics Problems And Solutions

**Viktor Korobov, Valery Ochkov**



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## **Kinetics Problems And Solutions** Book Review: Unveiling the Magic of Language

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