

Student Exploration: Calorimetry Lab

(Note: Gizmos can be done even on your phones! Contact Goff or Molik if you need your log-in information)

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.



Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.

1. On the SIMULATION pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the GRAPH tab, and click **Play** (▶).
 - A. What was the **Final temperature** of the copper and the water? 35 °C _____
 - B. How much did the temperature of the copper change? It went down 65 °C _____
 - C. How much did the temperature of the water change? _____ it went up 5 °C _____
2. Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Activity A: Heat transfer	<u>Get the Gizmo ready:</u> <ul style="list-style-type: none">• Click Reset (↺).	
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Question: What factors determine how heat energy transfers between objects?

1. **Predict:** In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g of 30.0 °C water.
 - A. How do you think increasing the water's mass would affect the final temperature?

Lab Activity Chemical Reaction Answer Key Calorimetry

Victor M. Corman



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Proceedings of the Workshop on Geochemical Modeling, 1986 Chemistry Nicholas R. Romano, Philip K. Cameron, 2002 Prepares students for the new standards and the commencement level PS Chemistry Test Challenges with content based multiple choice constructed response and real world thematic questions Stimulates skills based activities in reading writing and lab performance Correlates PS Chemistry key ideas and performance indicators on atomic concepts periodic table moles stoichiometry bonding behavior of matter kinetics organic chemistry oxidation reduction acids bases and salts nuclear chemistry Fosters mastery with practice on three recent tests *Chemistry, Key to Better Living* American Chemical Society, 1951 *An Assessment of the National Institute of Standards and Technology Measurement and Standards Laboratories* National Research Council, Division on Engineering and Physical Sciences, Board on Assessment of NIST Programs, 2002-10-26 This assessment of the technical quality and relevance of the programs of the Measurement and Standards Laboratories of the National Institute of Standards and Technology is the work of the 165 members of the National Research Council's NRC's Board on Assessment of NIST Programs and its panels These individuals were chosen by the NRC for their technical expertise their practical experience in running research programs and their knowledge of industry's needs in basic measurements and standards This assessment addresses the following The technical merit of the laboratory programs relative to the state of the art worldwide The effectiveness with which the laboratory programs are carried out and the results disseminated to their customers The relevance of the laboratory programs to the needs of their customers and The ability of the laboratories facilities equipment and human resources to enable the laboratories to fulfill their mission and meet their customers needs **Fossil Energy Update**, 1986 **Fiscal year 1986 Department of Energy authorization (basic research programs)** United States. Congress. House. Committee on Science and Technology. Subcommittee on Energy Development and Applications, 1985 Energy Research Abstracts, 1987 *U.S. Government Research & Development Reports*, 1969-10 **Summaries of FY ... Research in the Chemical Sciences**, 1985 *Comprehensive Supramolecular Chemistry II* George W. Gokel, Len Barbour, 2017-06-22 *Comprehensive Supramolecular Chemistry II* Second Edition Nine Volume Set is a one stop shop that covers supramolecular chemistry a field that originated from the work of researchers in organic inorganic and physical chemistry with some biological influence The original edition was structured to reflect in part the origin of the field However in the past two decades the field has changed a great deal as reflected in this new work that covers the general principles of supramolecular chemistry and molecular recognition experimental and computational methods in supramolecular chemistry supramolecular receptors dynamic supramolecular chemistry supramolecular engineering crystallographic engineered assemblies sensors imaging agents devices and the latest in nanotechnology Each section begins with an introduction by an expert in the field who offers an initial perspective on the development of the field Each article begins with outlining basic concepts before moving on to more advanced material

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