

LAB: SHAPES OF COVALENT MOLECULES & POLARITY

Introduction:

The most common chemical bond between two atoms is a **covalent bond**. The covalent bond consists of a pair of shared electrons, one from each atom. If this pair of electrons is shared between two atoms of equal electro negativities, the bond would be called a **nonpolar covalent bond**. However, in most cases, the pair of electrons is shifted toward the more electronegative element. A partial negative charge results on one side of the bond and a partial positive charge on the other. This type of covalent bond is called **polar covalent**.

Molecules composed of covalently bonded atoms may also be polar or nonpolar. For the molecule to be polar, it must, of course, have polar bonds. But the key factor for determining the polarity of a molecule is its shape. If the polar bonds (**dipoles**) are symmetrical around the central atom, they offset each other and the resulting molecule is nonpolar. However, if the dipoles are not symmetrical around the central atom, the electrons will be pulled to one end of this molecule and the resulting molecule is polar.

Objectives:

- (1) Predict each molecule's shape using your knowledge of hybridization.
- (2) Predict each molecule's polarity on the basis of its shape.

Equipment:

Paper and pencil

Procedure:

1. Prepare a data table according to the directions in the Analysis section below.
2. Complete your data table using the following compounds:

1. H_2	5. SO_2	9. CH_4	13. CH_2Cl
2. HBr	6. CO_2	10. HClO	14. HCOOH
3. H_2O	7. H_2CO	11. O_2	15. CO_3^{2-}
4. NH_3	8. C_2H_2	12. AlH_3	16. NH_4^+

Lab Shapes Of Covalent Molecules Answer Key

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