

Stoichiometry III – Theoretical Yield

Teacher

If we know what chemical reaction is occurring, we can use mole ratios to determine the limiting and excess reactants. We can do this by using the starting masses of each reactant to see how much of a product could be produced if all of the reactant was consumed. We then compare the number of moles produced by each reactant, and the smallest value is the theoretical yield – the amount of a product produced if all of the limiting reactant is consumed.

Practice Problems



a) Determine the limiting and excess reagents if you started with 14.8 g of C_3H_8 and 33.44 g of O_2 .

$$\frac{14.8 \text{ g } C_3H_8}{44.09 \text{ g } C_3H_8} \times \frac{1 \text{ mol } C_3H_8}{1 \text{ mol } C_3H_8} \times \frac{4 \text{ mol } H_2O}{1 \text{ mol } C_3H_8} \times \frac{18.02 \text{ g } H_2O}{1 \text{ mol } H_2O} = 24.2 \text{ g } H_2O$$

$$\frac{33.44 \text{ g } O_2}{32.00 \text{ g } O_2} \times \frac{1 \text{ mol } O_2}{5 \text{ mol } O_2} \times \frac{4 \text{ mol } H_2O}{1 \text{ mol } O_2} \times \frac{18.02 \text{ g } H_2O}{1 \text{ mol } H_2O} = 15.1 \text{ g } H_2O$$

Limiting: O_2
Excess: C_3H_8

b) What is the theoretical yield of H_2O . $15.1 \text{ g } H_2O$



a) Determine the limiting and excess reagents if 8.1 g of $Al_2(SO_4)_3$ is reacted with 8.1 g of $NaOH$

$$\frac{8.1 \text{ g } Al_2(SO_4)_3}{294.14 \text{ g } Al_2(SO_4)_3} \times \frac{1 \text{ mol } Al_2(SO_4)_3}{1 \text{ mol } Al_2(SO_4)_3} \times \frac{2 \text{ mol } Al(OH)_3}{1 \text{ mol } Al_2(SO_4)_3} \times \frac{78.00 \text{ g } Al(OH)_3}{1 \text{ mol } Al(OH)_3} = 4.3 \text{ g } Al(OH)_3$$

$$\frac{8.1 \text{ g } NaOH}{40.00 \text{ g } NaOH} \times \frac{1 \text{ mol } NaOH}{6 \text{ mol } NaOH} \times \frac{2 \text{ mol } Al(OH)_3}{1 \text{ mol } NaOH} \times \frac{78.00 \text{ g } Al(OH)_3}{1 \text{ mol } Al(OH)_3} = 5.3 \text{ g } Al(OH)_3$$

Limiting: $Al_2(SO_4)_3$
Excess: $NaOH$

b) What is the theoretical yield of aluminum hydroxide. $4.3 \text{ g } Al(OH)_3$



a) Determine the limiting and excess reagents if 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe .

$$\frac{25.4 \text{ g } Al_2O_3}{101.96 \text{ g } Al_2O_3} \times \frac{1 \text{ mol } Al_2O_3}{4 \text{ mol } Al_2O_3} \times \frac{3 \text{ mol } Fe_3O_4}{1 \text{ mol } Al_2O_3} \times \frac{231.55 \text{ g } Fe_3O_4}{1 \text{ mol } Fe_3O_4} = 43.3 \text{ g } Fe_3O_4$$

$$\frac{10.2 \text{ g } Fe}{55.85 \text{ g } Fe} \times \frac{1 \text{ mol } Fe}{9 \text{ mol } Fe} \times \frac{3 \text{ mol } Fe_3O_4}{1 \text{ mol } Fe} \times \frac{231.55 \text{ g } Fe_3O_4}{1 \text{ mol } Fe_3O_4} = 14.1 \text{ g } Fe_3O_4$$

Limiting: Fe
Excess: Al_2O_3

b) What is the theoretical yield of iron(II,III) oxide (aka magnetite).

$$14.1 \text{ g } Fe_3O_4$$

Chapter 3 Stoichiometry Exercises Answers

Katrin Zwirgmaier



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Table of Contents Chapter 3 Stoichiometry Exercises Answers

1. Understanding the eBook Chapter 3 Stoichiometry Exercises Answers
 - The Rise of Digital Reading Chapter 3 Stoichiometry Exercises Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 3 Stoichiometry Exercises Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 3 Stoichiometry Exercises Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 3 Stoichiometry Exercises Answers
 - Personalized Recommendations
 - Chapter 3 Stoichiometry Exercises Answers User Reviews and Ratings
 - Chapter 3 Stoichiometry Exercises Answers and Bestseller Lists
5. Accessing Chapter 3 Stoichiometry Exercises Answers Free and Paid eBooks
 - Chapter 3 Stoichiometry Exercises Answers Public Domain eBooks
 - Chapter 3 Stoichiometry Exercises Answers eBook Subscription Services
 - Chapter 3 Stoichiometry Exercises Answers Budget-Friendly Options
6. Navigating Chapter 3 Stoichiometry Exercises Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 3 Stoichiometry Exercises Answers Compatibility with Devices
 - Chapter 3 Stoichiometry Exercises Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 3 Stoichiometry Exercises Answers
 - Highlighting and Note-Taking Chapter 3 Stoichiometry Exercises Answers
 - Interactive Elements Chapter 3 Stoichiometry Exercises Answers

8. Staying Engaged with Chapter 3 Stoichiometry Exercises Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 3 Stoichiometry Exercises Answers
9. Balancing eBooks and Physical Books Chapter 3 Stoichiometry Exercises Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 3 Stoichiometry Exercises Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 3 Stoichiometry Exercises Answers
 - Setting Reading Goals Chapter 3 Stoichiometry Exercises Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 3 Stoichiometry Exercises Answers
 - Fact-Checking eBook Content of Chapter 3 Stoichiometry Exercises Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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