

# Chapter Test

Form G

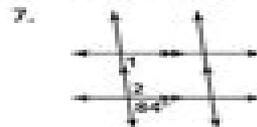
## Chapter 3

Decide whether each statement must be true or false. Use the figure for Exercises 1–8.

1.  $\angle 2$  and  $\angle 7$  are alternate interior angles.
2.  $\angle 5$  and  $\angle 6$  are same-side interior angles.
3.  $\angle 1$  and  $\angle 3$  are corresponding angles.
4. If  $\angle 14$  and  $\angle 15$  are congruent, then  $a \parallel b$ .
5. If  $\angle 10$  and  $\angle 11$  are supplementary, then  $c \parallel d$ .



Find  $m\angle 1$  and  $m\angle 2$ . Determine in each exercise whether  $\angle 1$  and  $\angle 2$  are alternate interior angles, same-side interior angles, or corresponding angles.



9. Graph the line  $y = -x + 1$ . Draw the line parallel to this line that contains  $(-1, -2)$ .
10. Graph the line  $y = -\frac{1}{2}x - 1$ . Draw the line perpendicular to this line that contains  $(2, 1)$ .

# Chapter Test Form G Geometry Test

**Chao Zhang**



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web exothermic reactions are chemical reactions that produce heat in section ii b of chapter 3 we defined the heat flow  $q$  as negative when heat flows from the system to the surroundings thus exothermic reactions have a negative  $\Delta h$  of reaction the word exothermic is derived from exo the greek word for outside and therme the greek word

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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example  $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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 temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable  
 measure an appropriate volume of each liquid eg 25 cm 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kj less  
 exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of  
 sodium hydroxide solution with hcl as 57.9 kJ mol<sup>-1</sup>

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 produce salts neutralisation is the reaction between an acid and a base

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic  
 eg 1  $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$  for which  $\Delta H = -57 \text{ kJ mol}^{-1}$  eg 2  $\text{HNO}_3(\text{aq}) + \text{KOH(aq)} \rightarrow \text{KNO}_3(\text{aq}) + \text{H}_2\text{O(l)}$  for which  $\Delta H = -57 \text{ kJ mol}^{-1}$

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example  $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o aq}$  you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts

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