

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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Table of Contents Chemistry Test Answers

1. Understanding the eBook Chemistry Test Answers
 - The Rise of Digital Reading Chemistry Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Test Answers
 - Personalized Recommendations

- Chemistry Test Answers User Reviews and Ratings
- Chemistry Test Answers and Bestseller Lists
- 5. Accessing Chemistry Test Answers Free and Paid eBooks
 - Chemistry Test Answers Public Domain eBooks
 - Chemistry Test Answers eBook Subscription Services
 - Chemistry Test Answers Budget-Friendly Options
- 6. Navigating Chemistry Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Test Answers Compatibility with Devices
 - Chemistry Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Test Answers
 - Highlighting and Note-Taking Chemistry Test Answers
 - Interactive Elements Chemistry Test Answers
- 8. Staying Engaged with Chemistry Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Test Answers
- 9. Balancing eBooks and Physical Books Chemistry Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Test Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Test Answers
 - Setting Reading Goals Chemistry Test Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Test Answers
 - Fact-Checking eBook Content of Chemistry Test Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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