

MULTIPLE CHOICE QUESTIONS

1. When potassium dichromate crystals are heated with hydrochloric acid, we get
(A) Oxygen (B) Hydrogen (C) Sulphur dioxide
(D) Hydrogen sulphide (E) Chlorine Ans : E
2. Which of the following elements has lowest melting point?
(A) Copper (B) Niobium (C) Molybdenum
(D) Tungsten (E) Mercury Ans : E
3. In acidic medium if potassium permanganate is used as an oxidising agent, we get
(A) MnO_4^{2-} (B) Mn^{2+} (C) MnO_2
(D) Mn_2O_7 (E) None of these Ans : B
4. Among the following outermost configurations of transition metals, which shows the highest oxidation state?
(A) $3d^3 4s^2$ (B) $3d^5 4s^1$ (C) $3d^4 4s^2$
(D) $3d^6 4s^2$ (E) $3d^7 4s^2$ Ans : C
5. Highest (+7) oxidation state is shown by
(A) Co (B) Cr (C) V
(D) Mn (E) Ni Ans : D
6. Potassium permanganate on heating gives
(A) K_2MnO_4, MnO_2, O_2 (B) MnO_2, O_2, KOH (C) K_2MnO_4, O_2

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