

# Oxidation–Reduction Titrations

Inquiry Guidance and AP® Chemistry Curriculum Alignment



## Introduction

Determining the amount of a particular substance in a sample or product is a common task in analytical chemistry. If the product contains a substance that can be oxidized, then it is possible to determine the number of moles of that substance by titrating the sample with a strong oxidizing agent. In this lab, an oxidizing solution will be standardized and then used to determine the number of moles of oxalic acid.

## Opportunities for Inquiry

Standardizing potassium permanganate by redox titration, then using that solution for quantitative analysis of another substance, fulfills key learning objectives involving redox titration, chemical reactions, and stoichiometry. Oxidation–reduction titrations require students to develop science practice skills for identifying data that must be collected and designing procedures to answer scientific questions.

The classic experiment can be modified to a guided-inquiry format by incorporating the following suggestions. Increasing the level of student preparation will also improve their level of understanding.

- Introduce the lab by demonstrating the general setup for redox titrations. Guide students to design the actual experimental procedure through a series of leading questions. What is the stoichiometry of the reaction involved in the redox titration? Determine a suitable concentration of potassium permanganate to analyze 0.25 M oxalic acid. What quantity of the primary standard, iron(II) ammonium sulfate, should be used in the standardization procedure? What is the purpose of adding an acid to the substance being titrated? Does the quantity of acid affect the stoichiometry of the redox titration? Answering these questions provides authentic opportunities for inquiry?
- Take away the data tables and post-lab questions. Replace worksheet calculations with a detailed overview of the experiment describing the general calculations. “The purpose of this lab is to standardize a solution of potassium permanganate by redox titration with a standard solution of iron(II) ions. A solution of oxalic acid is then titrated with the permanganate solution to determine the exact concentration of oxalic acid.”
- Extend the lab to analyze unknown and commercial products. Challenge students to determine the oxalic acid content in common household cleaning products by oxidation–reduction titration.
- Can students “make the leap” and generalize from one type of redox titration to another? Redox titrations using iodine are perhaps the most versatile analytical procedures for redox reactions. The amount of vitamin C, a strong reducing agent, is analyzed by titration with  $I_2$ . Ask students to research iodometric titration experiments and then design an experiment to analyze the vitamin C content in their favorite fruit juices.

## Alignment with AP Chemistry Curriculum Framework—Big Ideas 1 and 3

### Enduring Understandings and Essential Knowledge

All matter is made of atoms. There are a limited number of types of atoms; these are the elements.

(Enduring Understanding 1A)

1A): The mole is the fundamental unit for counting numbers of particles on the macroscopic level and allows quantitative connections to be drawn between laboratory experiments, which occur at the macroscopic level, and chemical processes, which occur at the atomic level.

Atoms are conserved in physical and chemical processes. (Enduring Understanding 1T)

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