

9/2/16

S. Sai Krishna

Date \_\_\_\_\_  
Page \_\_\_\_\_  
ClassmateThermochemistry

$$Q = \Delta E - W$$

$$Q = \Delta E - (-P\Delta V)$$

$$Q = \Delta E + P\Delta V$$

- At constant Volume.

$$Q = \Delta E$$

- At const. Pressure.

$$Q_p = \Delta E + P\Delta V$$

$$H = \Delta E + P\Delta V$$

$$H = \Delta E + \Delta nRT$$

$\Delta n =$  no. of moles of Prod  
- no. of moles of react.

→ For  $\Delta H$ .

- If  $\Delta H$  is -ve. reaction is exothermic.
- If  $\Delta H$  is +ve. reaction is endothermic.

Q. The heat of combustion of ethane at 17°C and const volume is  $-322.1 \text{ kJ}$ . Calculate the heat of combustion at constant pressure.



# Be Engineering Chemistry Notes 2016

**L Manion**



## **Be Engineering Chemistry Notes 2016:**

## **Be Engineering Chemistry Notes 2016** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Be Engineering Chemistry Notes 2016**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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