

Density is a Periodic Property

Discovering an Element

Introduction

"With the periodic and atomic relations now shown to exist between all atoms and the properties of their elements, we see the possibility not only of testing the absence of some of them but even of determining... properties of them as yet unknown elements."

Dmitri Mendeleev proposed the periodic law for the classification of elements in 1869–1871. After observing trends in the properties of elements when they were arranged in order of increasing atomic mass, Mendeleev made a startling prediction. He predicted the existence and properties of at least three undiscovered elements. Mendeleev saw what other scientists before him had missed—he saw what would come!

Concepts

- Periodic law
- Density
- Group IV elements
- Period number

Background

At the time Mendeleev proposed the periodic law, the foundation of the modern Periodic Table for the classification of elements, 63 elements were known. Their physical and chemical properties had been studied and their atomic masses measured. Mendeleev arranged the known elements in a calendar-like table of rows and columns in order of increasing atomic mass and repeating chemical properties. It is at this point, however, that Mendeleev made a giant leap of discovery—he suggested that there were some gaps or missing elements in the list of known elements.

Among the Group IV elements in Mendeleev's classification scheme, carbon appeared in the second row, followed by silicon in the third row. Both tin and lead shared similar chemical properties with carbon and silicon and were also known at this time. Because of their high atomic masses, however, these metals were placed in later rows of Mendeleev's Group IV column of elements. In 1871, Mendeleev proposed that there existed an as-yet-unknown element beneath silicon in the Group IV elements. He named the missing element *eka-silicon* and predicted its physical properties (atomic mass, melting point, density and specific heat). In 1886, the element germanium was discovered by the German chemist Clemens Winkler. In his report of the discovery, Winkler stated: "There can be no longer any doubt that the new element is no other than the *eka-silicon* postulated 15 years ago by Mendeleev."

Within 15 years of Mendeleev's prediction of the existence of missing elements, three of the elements had been discovered, their properties in excellent agreement with those predicted by Mendeleev. Is it possible to recreate some of the excitement that followed the prediction and discovery of Mendeleev's missing elements?

Experiment Overview

The purpose of this experiment is to measure mass and volume data for alkali, tin and lead; calculate their densities; and use these results to predict the density of germanium, Mendeleev's "undiscovered" element in the Group IV family of elements. The volume of the elements will be measured by water displacement.

Pre-Lab Questions

- One of the elements Mendeleev predicted was *eka-aluminum*, corresponding to a gap in the fourth row or period of the Group IIIA elements, between aluminum and indium. The density of aluminum (period 3) is 2.70, that of indium (period 5) 7.31 and that of thallium (period 6) 11.85 g/cm³. Make a graph of period number on the *x*-axis versus density on the *y*-axis for each of these elements.
- Use your graph to predict the density of *eka-aluminum*. What known element in the modern Periodic Table corresponds to *eka-aluminum*? Look up the density of the modern element in a reference source, and record its actual and predicted density values.
- How do the actual and predicted density values compare? Use the following equation to calculate the percent error between the predicted and actual values for the density of *eka-aluminum*.

$$\text{percent error} = \frac{|\text{actual} - \text{predicted}|}{\text{actual}} \times 100\%$$

Materials

- Lead shot, Pb, 35–40 g
- Silicon lumps, Si, 8–10 g
- Tin shot, Sn, 25–30 g
- Water
- Balance, centigrams (0.01 g precision)
- Beakers, 50-mL, or small cups, 3
- Tweezers or tongs
- Graduated cylinder, 25-mL
- Marking pencil or pen
- Paper towels

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Puja Mehta



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