

Class XII
Chemistry
Ch. 2: Solutions

Important formulae & Concepts

1. Mass percentage of a component (w/w)
$$= \frac{\text{Mass of component in solution}}{\text{Total mass of solution}} \times 100$$
- Volume percentage of a component (v/v)
2.
$$= \frac{\text{Volume of the component}}{\text{Total volume of solution}} \times 100$$
3. Mole fraction of a component (x) =
$$\frac{\text{Number of moles of the component}}{\text{Total number of moles of all components}}$$
4. Parts per million =
$$\frac{\text{Number of parts of component}}{\text{Total number of parts of all components of solution}} \times 10^6$$
5. Molarity =
$$\frac{\text{Number of moles of solute}}{\text{Volume of solution in litres}}$$
6. Molality =
$$\frac{\text{Number of moles of solute}}{\text{Mass of solvent in kilograms}}$$
7. Normality =
$$\frac{\text{Number of gram equivalent of solute}}{\text{Volume of solution in litres}}$$
8.
$$\frac{p_1^{\circ} - p_1}{p_1^{\circ}} = x_2$$
9.
$$\Delta T = T_b - T_b^{\circ}$$
$$\Delta T_b = \frac{K_b \times 1000 \times w_2}{M_2 \times w_1}$$
10.
$$\Delta T = T_f^{\circ} - T_f$$
$$\Delta T_f = \frac{K_f \times 1000 \times w_2}{M_2 \times w_1}$$

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DP Hallahan



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