

Experiment 25

Observations and Calculations: pH: Buffers and Their Properties

A. Determination of pH by the Use of Acid-Base Indicators

Indicator	Color with 0.1 M Solution of			
	HCl	NaH ₂ PO ₄	HClO ₂	ZnSO ₄
Methyl violet	blue	dark purple	dark purple	purple
Thymol blue	red	yellow	orange	yellow
Methyl yellow	red	yellow	red-orange	yellow
Congo red	purple	brown	dark purple	red-orange
Bromocresol green	yellow	green	yellow	green
pH range	1-3	2-4	2-3	2-5

Circle the observation(s) for each solution that was (were) most useful in estimating the pH range.

B. Measurement of the pH of Some Typical Solutions

Record the pH and the color observed with bromocresol green for each of the 0.1 M solutions that were tested.

	NaCl ¹	Na ₂ CO ₃ ²	NaC ₂ H ₃ O ₂ ³	NaHSO ₄ ⁴
pH	4.94	10.97	7.17	1.67
Color	green	blue	blue	yellow

For any two solutions having a pH less than 6 or greater than 8, write a net ionic equation to explain qualitatively why the solution has that pH.

Solution Na₂CO₃ Equation $\text{CO}_3^{2-} + \text{H}_2\text{O} \rightleftharpoons \text{HCO}_3^- + \text{OH}^-$

Solution NaHSO₄ Equation $\text{HSO}_4^- + \text{H}_2\text{O} \rightleftharpoons (\text{SO}_4)^{2-} + \text{H}_3\text{O}^+$

Explain why the color observed with bromocresol green for each of the four solutions is reasonable, given the pH.

C. Some Properties of Buffers

Buffer system selected HCO₃⁻ - CO₃²⁻ HB is hydrogen carbonate-carbonate (name the acid)

1. pH of buffer 9.28 [H⁺] 5.24×10^{-10} M pK_a (by Eq. 4b) _____

2. pH of diluted buffer 9.24 [H⁺] 5.75×10^{-10} M pK_a _____

pH after addition of 5 drops NaOH 9.34

pH after addition of 5 drops HCl 9.25

Comment on your observations in Parts 1 and 2.

(continued on following page)

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AW Chickering



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