

Physics

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(Chapter 12)(Atoms)
(Class 12)

The numerical value of the taken quantity will be:

$$\frac{1}{4\pi\epsilon_0} \times \frac{e^2}{m_e v^2}$$

$$= 9 \times 10^9 \times \frac{(1.6 \times 10^{-19})^2}{9.1 \times 10^{-31} \times (3 \times 10^6)^2}$$

$$= 2.81 \times 10^{-10} \text{ m}$$

Hence, the numerical value of the taken quantity is much smaller than the typical size of an atom.

(b) Charge on an electron, $e = 1.6 \times 10^{-19} \text{ C}$

Mass of an electron, $m_e = 9.1 \times 10^{-31} \text{ kg}$

Planck's constant, $h = 6.63 \times 10^{-34} \text{ Js}$

Let us take a quantity involving the given quantities as

$$\frac{4\pi\epsilon_0 \left(\frac{h}{2\pi}\right)^2}{m_e v^2}$$

Where,

$\epsilon_0 =$ Permittivity of free space and

$$\frac{1}{4\pi\epsilon_0} = 9 \times 10^9 \text{ N m}^2 \text{ C}^{-2}$$

The numerical value of the taken quantity will be:

$$4\pi\epsilon_0 \times \frac{\left(\frac{h}{2\pi}\right)^2}{m_e v^2}$$

$$= \frac{1}{9 \times 10^9} \times \frac{\left(\frac{6.63 \times 10^{-34}}{2 \times 3.14}\right)^2}{9.1 \times 10^{-31} \times (1.6 \times 10^{-10})^2}$$

$$= 0.53 \times 10^{-10} \text{ m}$$

Hence, the value of the quantity taken is of the order of the atomic size.



Question 12.15:

The total energy of an electron in the first excited state of the hydrogen atom is about -3.4 eV .

- What is the kinetic energy of the electron in this state?
- What is the potential energy of the electron in this state?
- Which of the answers above would change if the choice of the zero of potential energy is changed?

Answer 12.15:

(a) Total energy of the electron, $E = -3.4 \text{ eV}$

Kinetic energy of the electron is equal to the negative of the total energy.

$$\Rightarrow K = -E \quad \Rightarrow -(-3.4) = +3.4 \text{ eV}$$

Hence, the kinetic energy of the electron in the given state is $+3.4 \text{ eV}$.

(b) Potential energy (U) of the electron is equal to the negative of twice of its kinetic energy.

$$\Rightarrow U = -2K \quad \Rightarrow -2 \times 3.4 = -6.8 \text{ eV}$$

Hence, the potential energy of the electron in the given state is -6.8 eV .

(c) The potential energy of a system depends on the reference point taken. Here, the potential energy of the reference point is taken as zero. If the reference point is changed, then the value of the potential energy of the system also changes. Since total energy is the sum of kinetic and potential energies, total energy of the system will also change.

Question 12.16:

If Bohr's quantisation postulate (angular momentum = $nh/2\pi$) is a basic law of nature, it should be equally valid for the case of planetary motion also. Why then do we never speak of quantisation of orbits of planets around the sun?

Answer 12.16:

We never speak of quantisation of orbits of planets around the Sun because the angular momentum associated with planetary motion is largely relative to the value of Planck's constant (h). The angular momentum of the Earth in its orbit is of the order of $10^{35}h$. This leads to a very high value of quantum levels n of the order of 10^{35} . For large values of n , successive energies and angular momenta are relatively very small. Hence, the quantum levels for planetary motion are considered continuous.

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Table of Contents Physics Chapter 12 Answers

1. Understanding the eBook Physics Chapter 12 Answers
 - The Rise of Digital Reading Physics Chapter 12 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics Chapter 12 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Chapter 12 Answers
 - Personalized Recommendations
 - Physics Chapter 12 Answers User Reviews and Ratings
 - Physics Chapter 12 Answers and Bestseller Lists
5. Accessing Physics Chapter 12 Answers Free and Paid eBooks
 - Physics Chapter 12 Answers Public Domain eBooks
 - Physics Chapter 12 Answers eBook Subscription Services
 - Physics Chapter 12 Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Physics Chapter 12 Answers Compatibility with Devices
 - Physics Chapter 12 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physics Chapter 12 Answers
 - Highlighting and Note-Taking Physics Chapter 12 Answers
 - Interactive Elements Physics Chapter 12 Answers
8. Staying Engaged with Physics Chapter 12 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics Chapter 12 Answers
9. Balancing eBooks and Physical Books Physics Chapter 12 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics Chapter 12 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physics Chapter 12 Answers
 - Setting Reading Goals Physics Chapter 12 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physics Chapter 12 Answers
 - Fact-Checking eBook Content of Physics Chapter 12 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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