

**Unit 1****1A Personality****Exercise 1** (page 1)

- 1 shy 2 nasty 3 serious 4 stupid 5 weak 6 polite  
7 generous 8 talkative 9 patient 10 honest 11 arrogant

**Exercise 2** (page 1)

The opposite of hard-working is lazy.

**Exercise 3** (page 1)

- 1 dishonest 2 generous 3 rude 4 talkative 5 clever  
7 nasty 8 lazy 9 funny 10 impatient

**Exercise 4** (page 1) **LISTENING 2** (page 1)

- Terry: I love going out with my friends. Terry: He always pays for everybody. He just enjoys spending money! He's so different from his brother. That's my dad of course! The only problem with Harry is that he can be rather impolite. He isn't very polite and doesn't wait before he speaks. The other day he really upset my dad when he told him he was mean.
  - Jodie: My mum's amazing, really. She never loses her temper, she just stays really calm. She's got four children, three, my brother and my two sisters and that to ask who do things best or best at things, please we actually do them - but she doesn't get angry. She also works full time at the supermarket and doesn't get home till about 10. I don't know how she fits it all in. She never stops!
  - Martin: My sister and I have completely different personalities. I like meeting new people and am happy to chat to people I don't know - at parties, for example. But Sophie's a bit nervous in situations like that. She wants for people to come to her. But when she gets to know you and starts talking she really makes you laugh.
  - Fiona: We're twins, but we aren't identical. We don't look alike, and our personalities aren't alike, either. For example, I don't say very much, normally I'm fairly quiet. Louise is the complete opposite. She never stops! But she doesn't talk about herself. I don't know for example she got fantastic marks in her exams but you never hear her talking about it.
- 1 generous, rude 2 patient, hard-working  
3 shy, funny 4 talkative, modest

**Challenge!** (page 1) **LISTENING 2** (page 1)

- We can be rather impolite.
  - She never loses her temper, she just stays really calm.
  - Sophie's a bit nervous in situations like that.
  - I'm fairly quiet, Louise is the complete opposite.
- 1 rather impolite  
2 just, bit nervous  
4 fairly quiet, complete opposite

**Exercise 5** (page 1)

- 2 irresponsible 3 insensitive 4 unfair 5 irrational  
6 immature 7 incoherent 8 disloyal

**Exercise 6** (page 1)

- 1 sensitive 2 fit 3 immature 4 irrational 5 tolerant  
6 disloyal 7 fair 8 irresponsible

**1B Present simple and continuous****Exercise 1** (page 1)

- 2 ✓ 3 ✓ 4 I'm wearing trainers today. 5 ✓  
6 Look! That man is wearing pink shoes.  
7 'What does Sam do?' 'He's a dentist!'

**Exercise 2** (page 1)

- 1 Am, having 2 I'm shopping 3 are, looking for 4 never  
wear 5 like 6 know 7 do, need  
8 I'm going

**Exercise 3** (page 1)

- 1 a Am, coming 8 do, come  
2 a wear 5 I'm wearing  
3 a Don't, see 6 Am, taking  
4 a plays 9 is playing  
5 a hate 7 I'm enjoying

**Exercise 4** (page 1)

- 1 are, doing 2 I'm texting 3 he, going 4 are, laughing  
5 is going 6 are, doing 7 I'm going 8 isn't going

**Exercise 5** (page 1)

- 1 a 2 a 3 b 4 a 5 b 6 b 7 b 8 b

**Challenge!** (page 1)

Students' own answers

**1C Lifestyles****Exercise 1** (page 1)

Students' own answers

**Exercise 2** (page 1)

- 1 G 2 F 3 C 4 E 5 D 6 B 7 A

**Exercise 3** (page 1)

hospitable: friendly and welcoming  
fertile: good soil or land  
independent: free from outside control  
patriotic: supporting one's country  
sensible: based on good judgement  
vibrant: full of energy and life  
weaving: making a fabric  
woodcarving: making an object from wood  
ceramics: pots made from clay  
city dwellers: people who live in the city  
hanging out: spending time relaxing  
achievements: things done successfully  
prominent: important or famous

# Solutions Upper Intermediate Answer Key

**Hussin A.Rothana**



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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction  $2\text{Ag}^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s}$  label 4 2 3

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is  $2\text{Na}^+ \text{aq} + 2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + \text{SO}_4^{2-} \text{aq} \rightarrow 2\text{Na}^+ \text{aq} + \text{SO}_4^{2-} \text{aq} + \text{CuCl}_2 \text{s}$  the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation  $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{Na}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$  the spectator ions are  $\text{Cl}^-$  and  $\text{Na}^+$  net ionic equation  $\text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$   $\text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$  here is a video on spectator ions *chapter 8 04 ionic equations chemistry libretexts* - Mar 06 2023

web the six  $\text{NO}_3^- \text{aq}$  ions and the six  $\text{Na}^+ \text{aq}$  ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation  $3\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically  
present the starting materials end products

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chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change  
oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water h aq oh aq h  
2 o l the spectator ions in koh aq and hno 3 aq are k aq and no 3 aq how to identify and list the

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spectator ion can therefore be observed in the reaction of aqueous solutions of sodium carbonate and copper ii sulfate but  
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two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be  $2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} \rightarrow \text{Cu} \text{s} + 2\text{Cl}^- \text{aq}$  these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length

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