

## Quesiti e problemi

### 1 I nomi delle sostanze

1 Chi è stato il primo scienziato a tentare di razionalizzare i nomi delle sostanze chimiche?

Lavoisier

2 Quali nomenclature sono attualmente utilizzate?

nomenclatura tradizionale; nomenclatura di Stock;  
nomenclatura IUPAC

3 Che cos'è la IUPAC e di che cosa si occupa?

Associazione internazionale con lo scopo  
di razionalizzare la nomenclatura dei composti chimici.

### 2 Valenza e numero di ossidazione

4 Nel composto  $N_2O_5$ , gli atomi di azoto hanno valenza 5 e quelli di ossigeno hanno valenza 2.

► Qual è il numero di ossidazione dell'azoto?

+5

► Qual è il numero di ossidazione dell'ossigeno?

-2

5 Qual è il numero di ossidazione del silicio nel composto  $SiH_4$ ?

+4

6 Utilizzando la tabella 12.1, attribuisce il numero di ossidazione agli elementi presenti nelle seguenti formule costituite da due o tre elementi.

a)  $Ca(OH)_2$   $Ca: +2; O: -2; H: +1$

b)  $CaO$   $Ca: +2; O: -2$

c)  $MgSO_4$   $Mg: +2; O: -2; S: +6$

d)  $Fe_2O_3$   $Fe: +3; O: -2$

e)  $H_2O_2$   $H: +1; O: -1$

f)  $V_2O_5$   $V: +5; O: -2$

g)  $NH_3$   $N: -3; H: +1$

h)  $MnO_2$   $Mn: +4; O: -2$

i)  $HCO_3^-$   $H: +1; O: -2; C: +3$

l)  $PO_4^{3-}$   $P: +5; O: -2$

m)  $KMnO_4$   $K: +1; Mn: +7; O: -2$

n)  $PbI_2$   $Pb: +2; I: -1$

o)  $KNO_3$   $K: +1; N: +5; O: -2$

p)  $Cr_2O_3$   $Cr: +3; O: -2$

q)  $H_4SiO_4$   $H: +1; Si: +4; O: -2$

r)  $SrSO_4$   $Sr: +2; S: +6; O: -2$

s)  $P_2O_5$   $P: +5; O: -2$

t)  $MnO_4^{2-}$   $Mn: +6; O: -2$

u)  $Sb_2S_3$   $Sb: +3; S: -2$

v)  $IO_3^-$   $I: +5; O: -2$

7 Scrivi la formula di un composto dell'azoto con numero di ossidazione -3.

$NH_3$

8 Il manganese può avere numeri di ossidazione variabili tra +2 e +7. Qual è il suo n.o. nello ione manganato  $MnO_4^{2-}$ ?

+4

9 Utilizzando la tabella 12.1, attribuisce il numero di ossidazione agli elementi presenti nelle seguenti formule, costituite da uno o due elementi.

a)  $S_8$   $0$  b)  $H_2O$   $H: +1; O: -2$

c)  $Hg$   $0$  d)  $SO_3$   $S: +6; O: -2$

e)  $PH_3$   $P: -3; H: +1$  f)  $Cl_2O_7$   $Cl: +7; O: -2$

g)  $CH_4$   $C: -4; H: +1$  h)  $N_2$   $0$

i)  $LiH$   $Li: +1; H: -1$  l)  $CO$   $C: +2; O: -2$

m)  $Xe$   $0$  n)  $Al_2O_3$   $Al: +3; O: -2$

o)  $CuO$   $Cu: +2; O: -2$  p)  $Cu_2O$   $Cu: +1; O: -2$

q)  $AgCl$   $Ag: +1; Cl: -1$  r)  $Na_2S$   $Na: +1; S: -2$

s)  $F^-$   $-1$  t)  $SO_4^{2-}$   $S: +6; O: -2$

u)  $Ca^{2+}$   $+2$  v)  $MnO_4^-$   $Mn: +7; O: -2$

10 Trova i numeri di ossidazione degli elementi sottolineati.

a) Cl $_2O$   $+1$  b) Sg $H_2$   $-2$

c) Si $H_4$   $+4$  d) Ag $_2O$   $+1$

e) CO $_2$   $+4$  f) Si $Cl_4$   $+4$

g) K $_2SO_4$   $+1$  h) Ec $O$   $+2$

i) O $_2$   $0$  l) H $_2SO_3$   $+4$

m) H $NO_2$   $+3$  n) Ba $SO_4$   $+2$

o) H $_2S$   $-2$  p) NH $_3$   $-3$

11 Quali elementi formano ioni positivi (cationi) e quali formano ioni negativi (anioni)?

In generale: i metalli tendono a trasformarsi in cationi e i non metalli in anioni.

12 Trova i numeri di ossidazione degli elementi sottolineati.

a) Hg $^{2+}$   $+1$  b) Si $F_4^{2-}$   $+4$

c) S $_2O_8^{2-}$   $+7$  d) Sg $Cl_5^-$   $+4$

e) Gg $Cl_4^-$   $+4$  f) NH $_4^+$   $-3$

g) NO $_3^-$   $+5$  h) Cr $_2O_7^{2-}$   $+6$

i) Cl $O_4^-$   $+7$  l) NH $_2^+$   $-3$

m) NO $_3^-$   $+5$  n) Mg $O_4^{2-}$   $+6$

### 3 Leggere e scrivere le formule più semplici

13 Scrivi la formula bruta dei composti binari che si formano dai seguenti elementi.

a) K e I  $KI$  b) S (con n.o. -2) e Cl  $SO_2$

c) Sr e H  $SrH_2$  d) H e S  $H_2S$

# Soluzioni Esercizi Chimica Organica Zanichelli Hart

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