

## *Material Balance Calculations*

(Incorporating calculations of ideal gas laws, unit operations and chemical reactions etc.)

### THEORETICAL BACKGROUND

Ideal Gas Law :

$$PV = nRT$$

where,  $P$  = Total pressure of system  
 $V$  = Volume  
 $n$  = Number of moles  
 $T$  = Absolute temperature  
 $R$  = Universal gas constant

Values of universal gas constant ( $R$ ).

In different units, the values of universal gas constant are as under :

<i>Unit</i>	<i>Numerical value of R</i>
$\frac{(\text{atm.}) (\text{lit.})}{\text{gm mole } ^\circ\text{K}}$	0.08205
$\frac{(\text{atm.}) \text{m}^3}{\text{kg. mole } ^\circ\text{K}}$	0.08205
$\frac{(\text{atm.}) (\text{cc})}{\text{gm mole } ^\circ\text{K}}$	82.05
$\frac{(\text{Psi}) \text{ft}^3}{\text{lb. mole } ^\circ\text{R}}$	10.73
$\frac{(\text{atm.}) \text{ft}^3}{\text{lb. mole } ^\circ\text{R}}$	0.73
$\frac{(\text{mm Hg}) \text{lit.}}{\text{gm mole } ^\circ\text{K}}$	62.359
$\frac{\text{kcal}}{\text{kg. mole } ^\circ\text{K}}$	1.987

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**LL Leslie**



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## **Table of Contents Solved Examples In Chemical Engineering By Gk Roy Download**

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  - The Rise of Digital Reading Solved Examples In Chemical Engineering By Gk Roy Download
  - Advantages of eBooks Over Traditional Books
2. Identifying Solved Examples In Chemical Engineering By Gk Roy Download
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Solved Examples In Chemical Engineering By Gk Roy Download
  - User-Friendly Interface
4. Exploring eBook Recommendations from Solved Examples In Chemical Engineering By Gk Roy Download
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    - Interactive Elements Solved Examples In Chemical Engineering By Gk Roy Download
  8. Staying Engaged with Solved Examples In Chemical Engineering By Gk Roy Download
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Solved Examples In Chemical Engineering By Gk Roy Download
  9. Balancing eBooks and Physical Books Solved Examples In Chemical Engineering By Gk Roy Download
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Solved Examples In Chemical Engineering By Gk Roy Download
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Solved Examples In Chemical Engineering By Gk Roy Download
    - Setting Reading Goals Solved Examples In Chemical Engineering By Gk Roy Download
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Solved Examples In Chemical Engineering By Gk Roy Download
    - Fact-Checking eBook Content of Solved Examples In Chemical Engineering By Gk Roy Download
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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