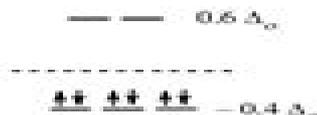


- b. $[\text{M}(\text{H}_2\text{O})_6]^{2+}$ M is second row transition metal, LFSE = $-2.4 \Delta_o$. This can be achieved with a low spin d^6 configuration. M^{2+} with 6 d electrons: M = Rh



- c. $[\text{MCl}_4]^-$ M is first row transition metal, 5 unpaired electrons. M^{2+} with 5 d electrons: Fe



- d. $\text{MCl}_2(\text{NH}_3)_2$ M is third row d^8 transition metal; two M–Cl stretching bands in IR. Second and third row d^8 complexes are often square planar. The presence of two M–Cl stretching bands implies *cis* geometry of the chloro ligands.

M^{2+} with 8 d electrons: M = Pt



- 10.4 Angular overlap calculations for d^8 and d^9 ions show no energy difference between D_{4h} and O_h when exclusively considering σ interactions. Both d^8 geometries have energies of $-3e_{\sigma}$; both d^9 geometries have energies of $-6e_{\sigma}$. In general, stability constants decrease as more ligands are added; the sequence for nickel is the common one. The huge drop in stability constant between the second and third ethylenediamine on Cu^{2+} is a result of the d^9 Jahn-Teller effect. The first two en ligands add in a square-planar geometry, with water molecules in the axial positions, and the coordination of these two monodentate ligands allows for the Jahn-Teller distortion. Adding a third en ligand requires a geometry change and with a preference for uniform M–N bond distances towards the six nitrogen atoms. This is counter to the Jahn-Teller distortion, and the third addition is much less favorable than the first two.

- 10.5 $[\text{M}(\text{H}_2\text{O})_6]^{2+}$ M is first row transition metal, $\mu = 3.9 \text{ BM}$. The magnetic moment implies 3 unpaired electrons, which would give rise to $\mu = \sqrt{3(3+5)} = 3.9 \text{ BM}$. There are two possibilities, d^3 and d^7 :



- 10.6
- | | | | |
|----|---|----|--|
| a. | $[\text{Cr}(\text{H}_2\text{O})_6]^{2+}$ $n = 4$
$\mu = \sqrt{4(6)} = 4.9 \mu_B$ | e. | $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ $n = 2$
$\mu = \sqrt{2(4)} = 2.8 \mu_B$ |
| b. | $[\text{Cr}(\text{CN})_6]^{3-}$ $n = 2$
$\mu = \sqrt{2(4)} = 2.8 \mu_B$ | f. | $[\text{Cu}(\text{en})_2(\text{H}_2\text{O})_2]^{2+}$ $n = 1$
$\mu = \sqrt{1(3)} = 1.7 \mu_B$ |
| c. | $[\text{FeCl}_4]^-$ $n = 5$
$\mu = \sqrt{5(7)} = 5.9 \mu_B$ | | |
| d. | $[\text{Fe}(\text{CN})_6]^{3-}$ $n = 1$ | | |

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