

## Titration Experiment 1: Acid-Base Standardization of NaOH Solution

Write chemical equation - check page 94 in your lab manual.



	2.0630 Titration 1	Titration 2	Titration 3	Titration 4	108 2.0717 Titration 5
Mass of KHP dispensed in beaker	<del>2.0630</del>	2.0566	1.9855	1.9993	<del>2.0717</del>
Volume NaOH used to titrate KHP unknown solution	29.03	29.13	28.12	28.32	28.50
Molarity of NaOH	0.345	0.346	0.3457	0.3456	0.3454
Average molarity of NaOH (avg ± standard deviation)	0.345				
% Precision					
Actual molarity NaOH unknown from Beyond Labz Unknown Tab					
% Error					

## Calculations

- Provide one representative calculation that you used to determine the molarity of NaOH. Be sure to show all units and use proper sig figs! We will assume you do this same calculation for all trials!

$$\frac{2.0630}{204.33} = 0.01009$$

$$\hookrightarrow \frac{0.01009}{0.0293} = 0.34438 \approx 0.345$$

# Acid Base Titration Chemistry If8766 Answer Key

**Bird**



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