

ACID-BASE TITRATION

Topic: Acid-Base

By determining the concentration of an acid (or base), we can determine the amount of base (or acid) of known concentration used in a titration experiment. The point at which indicators change color is known as the endpoint or color of the solution.

At the endpoint, the following equation



Solve the problems below.

1. A 25.0 mL sample of hydrochloric acid for the experiment with 10.0 mL of 0.100 M sodium hydroxide solution.

What is the molarity of the acid?
0.100 M

2. A 10.0 mL sample of 0.100 M acetic acid is neutralized by 10.0 mL of 0.100 M sodium hydroxide solution.

What is the molarity of the acid?
0.100 M

3. A 10.0 mL sample of 0.100 M acetic acid is neutralized by 10.0 mL of 0.100 M sodium hydroxide solution.

What is the molarity of the acid?
0.100 M

4. A 10.0 mL sample of 0.100 M acetic acid is neutralized by 10.0 mL of 0.100 M sodium hydroxide solution.

What is the molarity of the acid?
0.100 M

5. A 10.0 mL sample of 0.100 M acetic acid is neutralized by 10.0 mL of 0.100 M sodium hydroxide solution.

What is the molarity of the acid?
0.100 M

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