

Question 1: Kinetics of a Reaction and Equilibrium

1.1. Determining Reaction Order and Rate Law

- a. Second
- b. First
- c. Third
- d. Fourth
- e. Zeroth

1.2. Interpret the rate law for the reaction $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$ and determine the overall order of the reaction.

1.3. For the reaction $\text{A} + \text{B} \rightarrow \text{C}$, the rate law is $\text{rate} = k[\text{A}]^2[\text{B}]$. The partial orders are:

- a. 1
- b. 2
- c. 3
- d. 4
- e. 5
- f. 6
- g. 7
- h. 8
- i. 9
- j. 10

1.2. Determining Reaction Order and Rate Law

- a. $\text{rate} = k[\text{A}]^2[\text{B}]$
- b. $\text{rate} = k[\text{A}][\text{B}]$
- c. $\text{rate} = k[\text{A}][\text{B}]^2$

1.3. For the reaction $\text{A} + \text{B} \rightarrow \text{C}$, the rate law is $\text{rate} = k[\text{A}]^2[\text{B}]$. The partial orders are:

- a. 1
- b. 2
- c. 3
- d. 4
- e. 5
- f. 6
- g. 7
- h. 8
- i. 9
- j. 10

1.3. Determining Reaction Order and Rate Law

- a. $\text{rate} = k[\text{A}]^2[\text{B}]$

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