

Why Has the Value of Avogadro's Constant Changed Over Time?

William B. Jensen

Department of Chemistry, University of Cincinnati
Cincinnati, OH 45221-0172

Question

Why has the value of Avogadro's constant changed over time?

Ben Ruckberg
Department of Chemistry
University of Rhode Island
Kingston, RI 02881

Answer

As noted in an earlier column, the concept of Avogadro's constant or number (N_A) was not introduced by Avogadro in 1811, but rather by the French physical chemist, Jean Perrin (figure 1), in 1908 (1). It may be formally thought of as the value of the conversion factor between the gram (g) and the unified atomic mass unit (u), as may be seen from the requirement that the molecular mass of a given atom, molecule, or ion, as expressed in atomic mass units per entity, must be numerically equal to the molar mass of the same substance as expressed in grams per mole. Thus for example:

$$207.2 \text{ g/mol Pb} = (207.2 \text{ u/Pb atom})(1 \text{ g}/N_A \text{ u})(N_A \text{ Pb atoms/1 mol Pb})$$

There are two reasons why the value of N_A has changed over time. The first, and most obvious, is that any change in the standards used to define either the atomic mass unit or the gram will cause a shift in the value of their conversion factor and hence in the value of N_A . Such a change occurred in 1960 when the standard for the atomic mass unit was changed from the $O = 16$ scale to the $^{12}\text{C} = 12$ scale. A similar shift would have occurred around 1900 with the shift from the $H = 1$ to the $O = 16$ scale, but, of course, the concept of N_A was not a part of chemistry at that time (2). The second reason for a shift in the value of N_A has to do with an ever increasing ability to accurately measure the value of this constant, as illustrated by the selected examples given in the accompanying table (3).

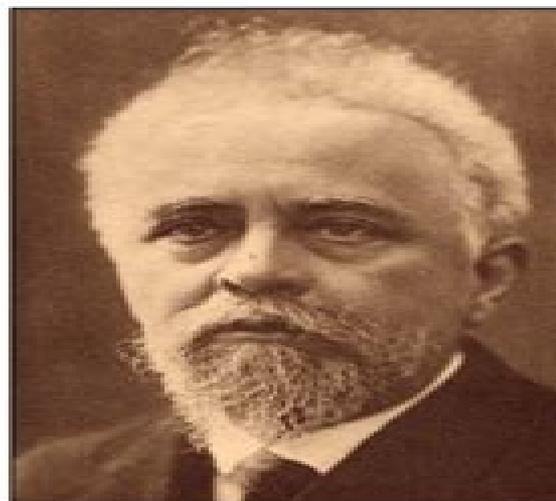


Figure 1. Jean Perrin (1870-1942).

The standard definition of N_A is that it is the same number of entities as there are carbon atoms in exactly 12 grams of carbon-12. Consequently students are often surprised when they are told that its numerical value is actually fixed by the conversion factor between the gram and the atomic mass unit. If chemists had failed to adopt the metric system and had instead continued to use some older conventional mass unit, such as the ounce, drachm or grain, to weigh chemicals in the laboratory, it would still have been expedient to maintain a numerical identity between atomic and molecular masses, as measured in atomic mass units, and the molar masses of the various substances as measured in the lab. Under such conditions, Avogadro's constant would have had a very different value, and it makes an interesting teaching exercise to assign

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