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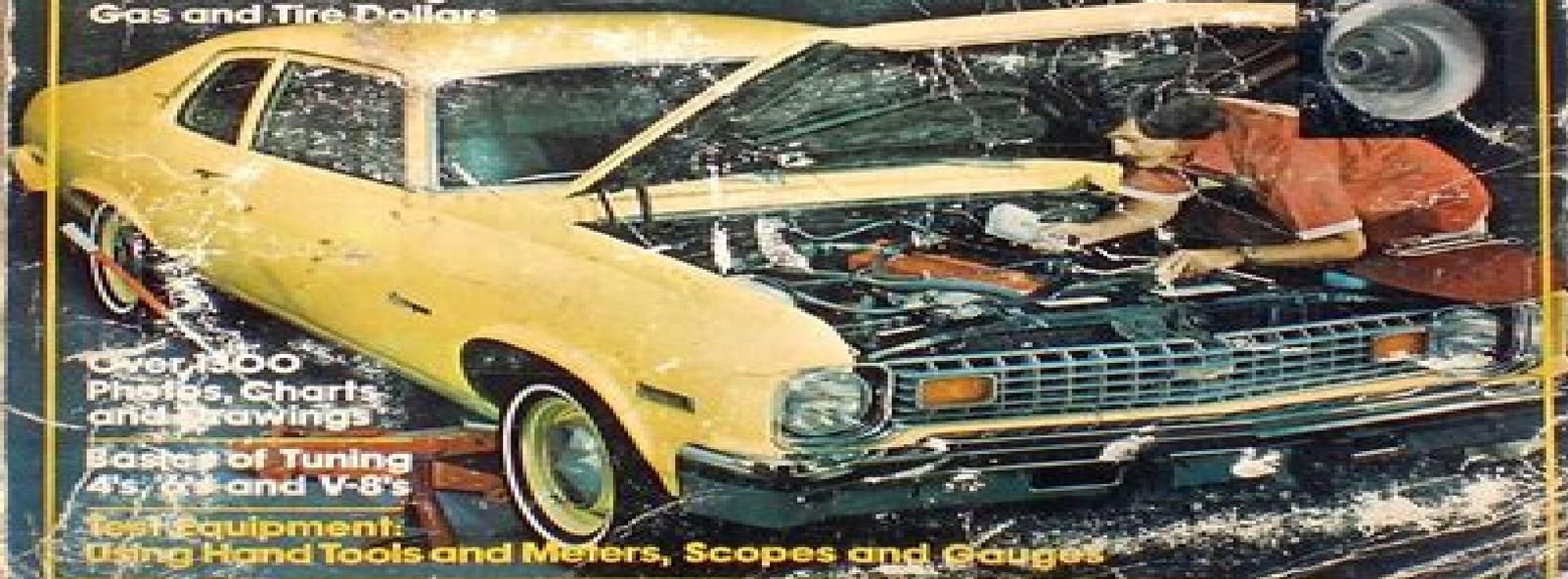
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# Auto Manual Book

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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web start with 8 3 mole h<sub>2</sub>s and multiply it with 2 mole h<sub>2</sub>o over 2 mole h<sub>2</sub>s from the equation 8 3 mole h<sub>2</sub>s x 2 mole h<sub>2</sub>o / 2 mole h<sub>2</sub>s x 18g h<sub>2</sub>o = 149 4g h<sub>2</sub>o is your theoretical yield so percent yield 137 1g h<sub>2</sub>o actual yield 149 4g h<sub>2</sub>o x 100 = 91 77 problems 6 8 are similar to problem 5

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calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160 g  $\text{Al}$   $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\text{yield actual} \times 100 = 260.0 \text{ g} \times 100 / 86.01 \text{ theoretical} = 302.3 \text{ g}$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield  $\text{actual yield} / 100$  based on this definition we would expect a percent yield

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which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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