

# Determination of Aspirin using Back Titration

This experiment is designed to illustrate techniques used in a typical *indirect* or *back* titration. You will use the NaOH you standardized last week to back titrate an aspirin solution and determine the concentration of aspirin in a typical analgesic tablet. You will be graded on your accuracy.

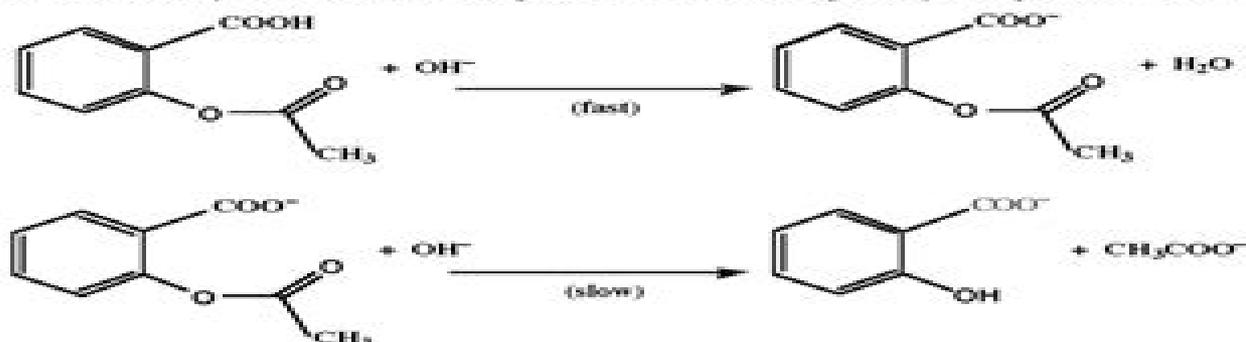
## Required Reading

D.C. Harris, *Quantitative Chemical Analysis* (6th ed., W. H. Freeman, NY, 2003) p. 130, 15.

## PreLab Quiz Topics

In addition to being able to explain the purpose of your experiment, the general procedure steps, the use of all chemicals in this experiment and any specific hazards, your prelab quiz may include explanations of any of the following terms: back titration, percent composition, the transition range of an indicator. You should be able to discuss the difference between a direct and a back titration, and know all relevant chemical reactions for this experiment. Describe the chemical and/or physical processes that will occur when you reach the endpoint of your titration. Explain clearly why we use a back titration for this particular analysis; i.e., why can't we just use the endpoint found from the initial NaOH titration?

Many reactions are slow or present unfavorable equilibria for direct titration. Aspirin is a weak acid that also undergoes slow hydrolysis; i.e., each aspirin molecule reacts with two hydroxide ions. To overcome this problem, a known **excess** amount of base is added to the sample solution and an HCl titration is carried out to determine the amount of unreacted base. This is subtracted from the initial amount of base to find the amount of base that actually reacted with the aspirin and hence the quantity of aspirin in the analyte.



## Chemicals and their Location

### Balance Room

Aspirin tablets (**Do not** bring your own sample)

### Above/on lab bench

Ethanol

Phenolphthalein, indicator soln

### In the Hood

Hydrochloric acid, conc., 37 wt %

### Your Drawer

Sodium hydroxide, standardized soln

# Analysis Of Aspirin Lab Report Conclusion

**AW Rasmussen**



## **Analysis Of Aspirin Lab Report Conclusion:**

**Exploring General Chemistry in the Laboratory** Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

**Forensics in Chemistry** Sara McCubbins, Angela Codron, 2012 Forensics seems to have the unique ability to maintain student interest and promote content learning I still have students approach me from past years and ask about the forensics case and specific characters from the story I have never had a student come back to me and comment on that unit with the multiple choice test at the end from the Introduction to Forensics in Chemistry The Murder of Kirsten K How did Kirsten K s body wind up at the bottom of a lake and what do wedding cake ingredients soil samples radioactive decay bone age blood stains bullet matching and drug lab evidence reveal about whodunit These mysteries are at the core of this teacher resource book which meets the unique needs of high school chemistry classes in a highly memorable way The book makes forensic evidence the foundation of a series of eight hands on week long labs As you weave the labs throughout the year and students solve the case the narrative provides vivid lessons in why chemistry concepts are relevant and how they connect All chapters include case information specific to each performance assessment and highlight the related national standards and chemistry content Chapters provide Teacher guides to help you set up Student performance assessments A suspect file to introduce the characters and new information about their relationships to the case Samples of student work that has been previously assessed and that serves as an answer key for you Grading rubrics Using Forensics in Chemistry as your guide you will gain the confidence to use inquiry based strategies and performance based assessments with a complex chemistry curriculum Your students may gain an interest in chemistry that rivals their fascination with Bones and CSI

**Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

**Pharmacology of Diabetes** Carl Erik Mogensen, Eberhard Standl, 1991 No detailed description available for Pharmacology of Diabetes

Ozark-Ouachita Mountains Vegetation Management (AR,OK) ,1990

**Cumulated Index Medicus** ,1990

Statistics by Example Terry Sincich, 1993

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Vegetation Management in the Ozark/Ouachita Mountains: Appendices ,1990

Vegetation Management in the Ozark/Ouachita Mountains: Comment letters and responses ,1990

**Danish Medical Bulletin** ,1954

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*Energy Research Abstracts* ,1981      **Science** John Michels (Journalist),2005 A weekly record of scientific progress  
*Digest of Comments on The Pharmacopœia of the United States of America and on the National Formulary ... Hygienic Laboratory (U.S.)*,1922      **Health 1998-1999** Richard Yarian,1998      **Nutrition, 98-99** Charlotte Cook-Fuller,Stephen Barrett,1998 This collection of articles reflect topics of interest and concern in the field of nutrition in the late 1990s  
Federal Practitioner for the Health Care Professionals of the VA, DoD, and PHS. ,1994

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