

ANSWERS TO END-OF-CHAPTER QUESTIONS

CHAPTER 8: ENERGY FROM ELECTRON TRANSFER

Emphasizing Essentials

- Define the terms *oxidation* and *reduction*.
 - Why must these processes take place together?

Answer:

- Oxidation is a process in which an atom, ion, or molecule *loses* one or more electrons. Reduction is a process in which an atom, ion, or molecule *gains* one or more electrons.
 - Electrons must be transferred from the species losing electrons to the species gaining electrons.
- Which of these half-reactions represent oxidation and which reduction? Explain your reasoning.
 - $\text{Fe}(s) \longrightarrow \text{Fe}^{2+}(aq) + 2 e^{-}$
 - $\text{Ni}^{4+}(aq) + 2 e^{-} \longrightarrow \text{Ni}^{2+}(aq)$
 - $2 \text{H}_2\text{O}(l) + 2 e^{-} \longrightarrow \text{H}_2(g) + 2 \text{OH}^{-}(aq)$

Answer:

- Oxidation. Iron loses two electrons to form the iron(II) ion.
 - Reduction. The nickel(IV) ion gains two electrons to form the nickel(II) ion.
 - Reduction. Water gains two electrons to the hydrogen ion and the hydroxide ion.
- You have seen several examples of oxidation–reduction reactions in this chapter. Now examine these equations and decide which are oxidation–reduction reactions and which are not. Explain your decisions.
 - $\text{Zn}(s) + 2 \text{MnO}_2(s) + \text{H}_2\text{O}(l) \longrightarrow \text{Zn}(\text{OH})_2(s) + \text{Mn}_2\text{O}_3(s)$
 - $\text{HCl}(aq) + \text{NaOH}(aq) \longrightarrow \text{NaCl}(aq) + \text{H}_2\text{O}(l)$
 - $\text{CH}_4(g) + 2 \text{O}_2(g) \longrightarrow \text{CO}_2(g) + 2 \text{H}_2\text{O}(g)$

Answer:

Parts **a** and **c** are redox reactions. Electrons must be transferred when an element (Zn in part **a**, O₂ in part **c**) reacts to form a compound. Part **b** is a neutralization reaction in which ions combine to form a soluble salt and water. Electron transfer does not occur in this case.

- Comment on the statement: Every combustion reaction is an oxidation–reduction reaction.

Answer:

This is a true statement. Combustion is the rapid combination of a fuel and oxygen to form products. In the case of burning a hydrocarbon, these products are CO₂ and H₂O. The oxygen is reduced, and the hydrogen and the carbon are oxidized to form H₂O and CO₂, respectively.

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