

KEY

Chemistry 12

Worksheet 1-3 - Reaction Mechanisms

1. It is known that compounds called *chlorofluorocarbons* (C.F.C.s) (eg. CFCl_3) will break up in the presence of ultraviolet radiation, such as found in the upper atmosphere, forming single chlorine atoms:



The Cl atoms then react with Ozone (O_3) as outlined in the following mechanism.



- a) Write the equation for the *overall reaction*. (Using steps 1 and 2)



- b) What is the *catalyst* in this reaction?



- c) Identify an *intermediate* in this reaction



- d) Explain how a *small* amount of chlorofluorocarbons can destroy a *large* amount of ozone.

When a Cl atom destroys an O_3 molecule (step 1), it is regenerated (step 2) so it can destroy more

- e) What breaks the bond in the CFCl_3 and releases the free Cl atom?

a photon ($h\nu$) of ultraviolet radiation

2. Given the following mechanism, answer the questions below:



- a) Give the equation for the *overall reaction*.



- b) What could the *catalyst* be in this mechanism?



- c) What is an *intermediate* in this mechanism?



Answers To Unit 5 Post Test Chemistry

G Orfield



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