

HOLT MODERN CHEMISTRY -- CHAPTER 3 REVIEW

The following pages contain the bulk (but not all) of the information for the chapter 3 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and **NOTHING** else, you should at least be able to **PASS** the test. ***** Test items will be recall, examples, and/or application of this content. *****

OUTCOMES

- 3.1: Explain law of conservation of mass (*T & R*)
- 3.1-3.2: Explain how the atom changed over time (Rutherford - Bohr models) (*F & PK*)
- 3.3: Contrast the nature and properties of the three subatomic particles: proton, neutron, and electron (*T & R*)
- 3.3: Use atomic number and mass number to find the number of subatomic particles of any given atom (*F & PK*)
- 3.3: Draw contrasting Bohr models of isotopes (*F & PK*)
- 3.3: Define mole and Avogadro's number (*T & R*)

3.1 THE ATOM: FROM PHILOSOPHICAL IDEA TO SCIENTIFIC THEORY

- **Vocabulary**
 - **law of conservation of mass** -- the law that states that mass cannot be created or destroyed in ordinary chemical and physical changes
 - **law of definite proportions** -- the law that states that a chemical compound always contains the same elements in exactly the same proportions by weight or mass
 - **law of multiple proportions** -- the law that states that when two elements combine to form two or more compounds, the mass of one element that combines with a given mass of the other is in the ratio of small whole numbers
- **Chapter Highlights**
 - The idea of atoms has been around since the time of the ancient Greeks. In the nineteenth century, John Dalton proposed a scientific theory of atoms that can still be used to explain properties of most chemicals today.
 - Matter and its mass cannot be created or destroyed in chemical reactions.
 - The mass ratios of the elements that make up a given compound are always the same, regardless of how much of the compound there is or how it was formed.
 - If two or more different compounds are composed of the same two elements, then the ratio of the masses of the second element combined with a certain mass of the first element can be expressed as a ratio of small whole numbers.

3.2 THE STRUCTURE OF THE ATOM

- **Vocabulary**
 - **atom** -- the smallest unit of an element that maintains the chemical properties of that element
 - **nuclear forces** -- the interaction that binds protons and neutrons, protons and protons, and neutrons and neutrons together in a nucleus
- **Chapter Highlights**
 - Cathode-ray tubes supplied evidence of the existence of electrons, which are negatively charged subatomic particles that have relatively little mass.
 - Rutherford found evidence for the existence of the atomic nucleus by bombarding gold foil with a beam of positively charged particles.
 - Atomic nuclei are composed of protons, which have an electric charge of +1, and (in all but one case) neutrons, which have no electric charge.
 - Atomic nuclei have radii of about 0.001 pm (pm = picometers; $1 \text{ pm} \times 10^{-12} \text{ m}$), and atoms have radii of about 40^{-270} pm.

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Camilla Rothe



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