

15% solution by mass means

Mass of solution = 100g

Mass of solute = 15g

Number of moles of $\text{H}_2\text{SO}_4 = 15/98$

H_2SO_4 Density = $\frac{\text{Mass}}{\text{Volume}}$

$1 \times 2 + 32 + 16 \times 4 = 98$ $1.2 = \frac{100}{\text{Volume}}$

$$\text{Volume} = \frac{100}{2} \times 10 = \frac{1000}{12} = 0.012 \text{ ml}$$

$\frac{100}{1.2}$ solution contain = $\frac{15}{98}$

1 solution contain = $\frac{15}{98} \times \frac{10}{100}$

1000 solution contain = $\frac{15}{98} \times \frac{1.2}{100} \times 1000$
= 1.8 M.

2.4 | MOLALITY

Number of moles of the solute present in 1000 g solvent. It is represented by 'M'.

2.5 | NORMALITY

Number of gram equivalents of a solute in 1000 ml solution. It is represented by 'N'.

Number of gram equivalents = $\frac{\text{Mass in gram}}{\text{equivalent Mass}}$

FORMULAE FOR EQUIVALENT MASS

Equivalent mass of acid = $\frac{\text{Molecular Mass of Acid}}{\text{Basicity}}$

Equivalent mass of base = $\frac{\text{Molecular Mass of Base}}{\text{Acidity}}$

Equivalent mass of salt = $\frac{\text{Formula mass of salt}}{\text{tvevalency of cation}}$

2.6 | MOLE FRACTION

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