

FORMULAS FOR IONIC COMPOUNDS

1. Write the chemical formula for the ionic compounds made from the following elements.

Elements	metal ion	non-metal ion	Chemical Formula	Name
potassium and iodine	K^{+1}	I^{-1}	KI	Potassium Iodide
barium and chlorine	Ba^{2+}	Cl^{-1}	$BaCl_2$	Barium Chloride
sodium and oxygen	Na^{+1}	O^{2-}	Na_2O	sodium Oxide
aluminum and nitrogen	Al^{3+}	N^{3-}	AlN	Aluminum Nitride
lithium and sulfur	Li^{+1}	S^{2-}	Li_2S	Lithium Sulfide
magnesium and oxygen	Mg^{+2}	O^{2-}	MgO	Magnesium Oxide
calcium and chlorine	Ca^{2+}	Cl^{-1}	$CaCl_2$	calcium chloride
boron and chlorine	B^{3+}	Cl^{-1}	BCl_3	boron chloride
beryllium and sulfur	Be^{2+}	S^{2-}	BeS	beryllium sulfide

2. Complete the chart below and write the ionic formulas to complete it. A few have been done for you.

	F	Cl	O^{2-}	S^{2-}	N^{3-}	P^{3-}
Li^{+}	LiF	$LiCl$	Li_2O	Li_2S	Li_3N	Li_3P
Na^{+}	NaF	$NaCl$	Na_2O	Na_2S	Na_3N	Na_3P
K^{+}	KF	KCl	K_2O	K_2S	K_3N	K_3P
Be^{2+}	BeF_2	$BeCl_2$	BeO	BeS	Be_3N_2	Be_3P_2
Mg^{2+}	MgF_2	$MgCl_2$	MgO	MgS	Mg_3N_2	Mg_3P_2
Ca^{2+}	CaF_2	$CaCl_2$	CaO	CaS	Ca_3N_2	Ca_3P_2
Al^{3+}	AlF_3	$AlCl_3$	Al_2O_3	Al_2S_3	AlN	AlP
Ga^{3+}	GaF_3	$GaCl_3$	Ga_2O_3	Ga_2S_3	GaN	GaP
B^{3+}	BF_3	BCl_3	B_2O_3	B_2S_3	BN	BP

Ionic Formula Writing Kit Answers

KJ Lindholm-Leary



Ionic Formula Writing Kit Answers:

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Questions set at the examinations College Entrance Examination Board, 1923 Questions Set at the Examinations, 1922 Questions Set at the Comprehensive Examinations of ..., 1924 *Chemistry at a Glance* Roger Owen, Sue King, 2005-06-27 This book aims to cover the specifications of the main examination boards for GCSE Double Science GCSE Single Science and the core content of GCSE Chemistry Where relevant Key Stage 3 material is summarized as an introduction to GCSE topics This serves as revision of work done prior to Key Stage 4 and a foundation for GCSE studies The book is **Chemistry** James E. Brady, Fred Senese, 2004-02-04 Publisher Description **Chemistry** Neil D. Jespersen, Alison Hyslop, 2021-11-02 *Chemistry The Molecular Nature of Matter* 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order Prentice Hall Chemistry, 2000 **Chemistry, an Introduction** Stanley M. Cherim, Leo E. Kallan, 1980 Basic Chemistry Steven S. Zumdahl, 2004 Description Not Yet Available **Prentice Hall Science Explorer** Michael J. Padilla, 2002 **Numerical Chemistry for Competitions** Anu Sharma, 2009 An ideal book for the students of XI and XII CBSE ISC and the State Boards who are using Core Curriculum and also useful for the students preparing for various Engineering Medical Entrance Examinations **Science** John Michels (Journalist), 1903 Vols for 1911 13 contain the Proceedings of the Helminothological Society of Washington ISSN 0018 0120 1st 15th meeting *Lutheran*

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