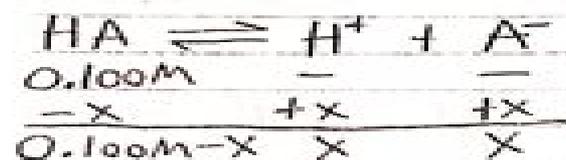


• Calculate K_a for the following acids:

1) 100.0 mL of 0.100 M HA with a pH of 3.40

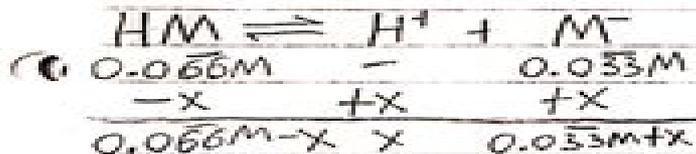


$$\text{pH} = 3.40 \rightarrow \text{pH} = -\log[\text{H}^+] \\
 3.40 = -\log(x) \rightarrow x = 3.98 \times 10^{-4}$$

$$K_a = \left(\frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \right) = \left(\frac{(x)^2}{(0.100\text{M}-x)} \right)$$

$$K_a = \left(\frac{(3.98 \times 10^{-4})^2}{(0.100\text{M} - 3.98 \times 10^{-4})} \right) \rightarrow \boxed{K_a = 1.59 \times 10^{-6}}$$

2) 100.0 mL of 0.100 M HM is mixed with 50.0 mL of 0.100 M NaM and has a pH of 5.50



$$\text{pH} = 5.50 \rightarrow \text{pH} = -\log[\text{H}^+] \\
 5.50 = -\log(x) \rightarrow x = 3.16 \times 10^{-6}$$

Dilution problem due to HM and NaM

$$(C_1)(V_1) = (C_2)(V_2) \\
 \rightarrow (0.100\text{M HM})(100.0\text{mL}) = (C_2)(150.0\text{mL}) \\
 (C_2) = 0.066\text{M HM}$$

Dilution problem due to NaM and HM

$$(C_1)(V_1) = (C_2)(V_2) \\
 (0.100\text{M NaM})(50.0\text{mL}) = (C_2)(150.0\text{mL}) \\
 (C_2) = 0.033\text{M NaM}$$

$$K_a = \left(\frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \right) = \left(\frac{(x)(0.033\text{M}+x)}{(0.066\text{M}-x)} \right)$$

$$K_a = \left(\frac{(3.16 \times 10^{-6})(0.033\text{M} + 3.16 \times 10^{-6})}{(0.066\text{M} - 3.16 \times 10^{-6})} \right) \rightarrow K_a = 1.58 \times 10^{-6}$$

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Rachel Sandford



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Table of Contents College Chemistry Problems And Solutions Pdf

1. Understanding the eBook College Chemistry Problems And Solutions Pdf
 - The Rise of Digital Reading College Chemistry Problems And Solutions Pdf
 - Advantages of eBooks Over Traditional Books
2. Identifying College Chemistry Problems And Solutions Pdf
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an College Chemistry Problems And Solutions Pdf
 - User-Friendly Interface

4. Exploring eBook Recommendations from College Chemistry Problems And Solutions Pdf
 - Personalized Recommendations
 - College Chemistry Problems And Solutions Pdf User Reviews and Ratings
 - College Chemistry Problems And Solutions Pdf and Bestseller Lists
5. Accessing College Chemistry Problems And Solutions Pdf Free and Paid eBooks
 - College Chemistry Problems And Solutions Pdf Public Domain eBooks
 - College Chemistry Problems And Solutions Pdf eBook Subscription Services
 - College Chemistry Problems And Solutions Pdf Budget-Friendly Options
6. Navigating College Chemistry Problems And Solutions Pdf eBook Formats
 - ePub, PDF, MOBI, and More
 - College Chemistry Problems And Solutions Pdf Compatibility with Devices
 - College Chemistry Problems And Solutions Pdf Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of College Chemistry Problems And Solutions Pdf
 - Highlighting and Note-Taking College Chemistry Problems And Solutions Pdf
 - Interactive Elements College Chemistry Problems And Solutions Pdf
8. Staying Engaged with College Chemistry Problems And Solutions Pdf
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers College Chemistry Problems And Solutions Pdf
9. Balancing eBooks and Physical Books College Chemistry Problems And Solutions Pdf
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection College Chemistry Problems And Solutions Pdf
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine College Chemistry Problems And Solutions Pdf
 - Setting Reading Goals College Chemistry Problems And Solutions Pdf
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of College Chemistry Problems And Solutions Pdf
 - Fact-Checking eBook Content of College Chemistry Problems And Solutions Pdf
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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