

**Standard-State Thermodynamic Values at 298.15 K:  
Enthalpy of Formation ( $\Delta H_f^\circ$ ), Free Energy of Formation ( $\Delta G_f^\circ$ ), and  
Absolute Entropy ( $S^\circ$ )**

Substance	$\Delta H_f^\circ$ (kJ/mol, $\Delta H_f^\circ$ )	$\Delta G_f^\circ$ (kJ/mol, $\Delta G_f^\circ$ )	$S^\circ$ (J/mol·K, $S^\circ$ )
<b>Aluminum</b>			
Al(s)	0	0	28.3
Al <sup>3+</sup> (aq)	-1675.7	-1609.3	82.9
Al(OH) <sub>3</sub> (s)	-1675.7	-1582.3	101.9
Al(OH) <sub>3</sub> (aq)	-1675.7	-1582.3	118.67
<b>Barium</b>			
Ba(s)	0	0	62.6
Ba <sup>2+</sup> (aq)	-855.26	-762.97	99.6
BaCO <sub>3</sub> (s)	-1218.8	-1012.9	112.1
BaCO <sub>3</sub> (aq) (BaCO <sub>3</sub> (aq))	-1218.8	-1012.9	219.08
BaCl <sub>2</sub> (s)	-858.64	-801.21	112.6
BaCl <sub>2</sub> (aq)	-858.64	-801.21	192.0
Ba(OH) <sub>2</sub> ·8H <sub>2</sub> O(s)	-3302.67	-3026.94	327.4
<b>Beryllium</b>			
Be(s)	0	0	14.6
Be <sup>2+</sup> (aq)	-112.8	-113.75	12.9
BeCO <sub>3</sub> (s)	-699.6	-596.3	126.7
BeCl <sub>2</sub> (s)	-499.8	-437.01	82.04
<b>Bismuth</b>			
Bi(s)	0	0	96.24
Bi <sub>2</sub> O <sub>3</sub> (s)	-513.04	-463.7	112.9
BiCl <sub>3</sub> (s)	-318.4	-273.34	117.0
<b>Boron</b>			
B(s)	0	0	6.0
B <sub>2</sub> O <sub>3</sub> (s)	-1273.5	-1203.7	202.4
<b>Bromine</b>			
Br <sub>2</sub> (l)	0	0	152.23
Br <sub>2</sub> (g)	30.91	7.84	245.46
Br <sup>-</sup> (aq)	-121.7	-109.75	82.4
HBr(g)	-36.4	-23.46	198.69
<b>Carbon</b>			
C(s)	0	0	5.7
C(s) (graph)	0	0	5.7
C(s) (diamond)	1.895	2.866	2.37
C(s) (amorph)	0	0	5.7
C(s) (graphite)	0	0	5.7
C(s) (fullerene)	1.895	2.866	2.37
C(s) (nanotube)	1.895	2.866	2.37

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**DP Hallahan**



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