

## Stoichiometry Lab: Vinegar and Baking Soda Do Not Write On This Sheet

**Purpose:** To predict the amount of Carbon Dioxide gas that should be produced in a chemical reaction; then calculate the % yield.



**Materials:** Baking Soda ( $\text{NaHCO}_3$ ), Vinegar ( $\text{CH}_3\text{COOH}$ ), and 2 plastic cups, scale.

**Procedure:** Find and record the mass of cup A. With cup A still on the scale, add approximately 10.0 g of baking soda to the cup. **The mass does not have to be exact, as long as you carefully record your results.**

Place cup B on the scale, weigh and record approximately 50.0 g of vinegar. **The mass does not have to be exact, as long as you carefully record your results.**

Slowly add vinegar to cup A until the reaction has stopped. **DO NOT** add all of the vinegar, just enough to complete the reaction. Reweigh and record both cup A and B. Calculate the mass of  $\text{CO}_2$  that escaped.

**Data:**

- a. Mass of Cup A \_\_\_\_\_ g
- b. Mass of Cup A and baking soda \_\_\_\_\_ g
- c. Calculate mass of baking soda (b - a) \_\_\_\_\_ g
- d. Mass of Cup B with vinegar \_\_\_\_\_ g
- e. Mass of Cup B after reaction \_\_\_\_\_ g
- f. Calculate mass of vinegar poured into Cup A (d - e) \_\_\_\_\_ g
- g. Mass of Cup A after reaction \_\_\_\_\_ g
- h. Calculate mass of product after reaction (g - a) \_\_\_\_\_ g
- i. Calculate baking soda + vinegar (c + f) \_\_\_\_\_ g
- j. Calculate mass of  $\text{CO}_2$  lost (i - h) \_\_\_\_\_ g

# Stoichiometry Lab Vinegar And Baking Soda Answers

**Sydney Young**



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