

NAME \_\_\_\_\_ DATE \_\_\_\_\_ CLASS \_\_\_\_\_

**CHAPTER 12** **STUDY GUIDE FOR CONTENT MASTERY**

**Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and its chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of  $\text{O}_2$ ,  $\text{N}_2$ , and  $\text{N}_2\text{O}_2$  are 32.00 g/mol, 28.02 g/mol, and 68.02 g/mol, respectively.

Amount of $\text{O}_2$	Amount of $\text{N}_2$	Amount of $\text{N}_2\text{O}_2$	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	NO	2 molecules $\text{O}_2$
2 molecules	4 molecules	4 molecules	2. $\text{O}_2$	4. 4 molecules NO
1.00 mol	2.00 mol	4. 2.00 mol	4. none	4. none
4.00 mol	4.00 mol	7. 4.00 mol	6. $\text{NO}$	6. 2.00 mol $\text{O}_2$
0.00 mol	1.00 mol	10. 7.00 mol	11. $\text{NO}$	10. 1.00 mol $\text{O}_2$
1.00 mol	0.50 mol	13. 2.00 mol	14. $\text{O}_2$	13. 2.00 mol NO
0.200 mol	0.200 mol	16. 0.200 mol	17. $\text{NO}$	16. 0.400 mol $\text{O}_2$
32.00 g	56.04 g	19. 92.02 g	20. none	21. none
64.00 g	56.04 g	22. 68.02 g	22. $\text{O}_2$	24. 30.12 g NO
64.00 g	28.02 g	25. 28.76 g	26. $\text{O}_2$	27. 1.24 g NO

balanced equation mole ratio = 2 mol  $\text{NO}$ /1 mol  $\text{O}_2$

16.00 g  $\text{O}_2$  = 1 mol  $\text{O}_2$ /32.00 g  $\text{O}_2$  = 0.5125 mol  $\text{O}_2$

20.00 g  $\text{NO}$  = 1 mol  $\text{NO}$ /30.01 g  $\text{NO}$  = 0.6664 mol  $\text{NO}$

actual mole ratio = 0.6664 mol  $\text{NO}$ /0.5125 mol  $\text{O}_2$  = 2.132 mol  $\text{NO}$ /1.000 mol  $\text{O}_2$

Because the actual mole ratio of  $\text{NO}$ / $\text{O}_2$  is larger than the balanced equation mole ratio of  $\text{NO}$ / $\text{O}_2$ , there is an excess of  $\text{NO}$ ;  $\text{O}_2$  is the limiting reactant.

Mass of  $\text{NO}$  used = 0.5125 mol  $\text{O}_2$  = 2 mol  $\text{NO}$ /1 mol  $\text{O}_2$  = 0.4250 mol  $\text{NO}$

0.4250 mol  $\text{NO}$  = 30.01 g  $\text{NO}$ /1 mol  $\text{NO}$  = 12.76 g  $\text{NO}$

Mass of  $\text{NO}_2$  produced = 0.4250 mol  $\text{NO}_2$  = 68.02 g  $\text{NO}_2$ /1 mol  $\text{NO}_2$  = 28.76 g  $\text{NO}_2$

Excess  $\text{NO}$  = 20.00 g  $\text{NO}$  = 12.76 g  $\text{NO}$  = 7.24 g  $\text{NO}$

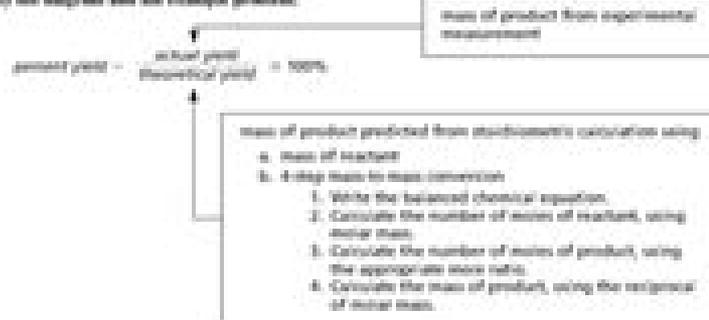
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**CHAPTER 12** **STUDY GUIDE FOR CONTENT MASTERY**

**Section 12.4 Percent Yield**

In your textbook, read about the percent of products.

Study the diagram and the example problem.



**Example Problem:** The following chemical equation represents the production of yellow oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and  $\text{Ga}_2\text{O}_3$  are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g  $\text{Ga}_2\text{O}_3$
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: 7.00 g Ga =  $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$  7.00 g Ga = 1 mol Ga/69.72 g Ga
- number of moles of product: 0.100 mol Ga =  $\frac{2 \text{ mol Ga}_2\text{O}_3}{1 \text{ mol Ga}}$  0.100 mol Ga = 2 mol  $\text{Ga}_2\text{O}_3$ /4 mol Ga
- theoretical yield: 0.200 mol  $\text{Ga}_2\text{O}_3$  =  $\frac{187.44 \text{ g Ga}_2\text{O}_3}{1 \text{ mol Ga}_2\text{O}_3}$  correct
- percent yield:  $\frac{7.42 \text{ g Ga}_2\text{O}_3}{187.44 \text{ g Ga}_2\text{O}_3} \times 100$  7.42 g  $\text{Ga}_2\text{O}_3$ /187.44 g  $\text{Ga}_2\text{O}_3$  = 100

# Standardized Test Practice Chapter 12 Chemistry Answers

**Y Pai**



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web oct 30 2023 two handfuls of blueberries or raspberries one heaped tablespoon of dried fruit such as sultanas currants or raisins three heaped tablespoons of vegetables such as sweetcorn or peas three heaped tablespoons of any pulse beans peas or lentils one cereal bowl of raw leafy greens such as lettuce watercress or spinach

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