

16. ^{137}I has a half-life of 8.04 days. Assuming you start with a 1.25 mg sample of ^{137}I , how much will remain after 13.8 days?

- A) 0.625 mg
 B) 0.268 mg
 C) 0.422 mg
 D) 0.388 mg

$$k = \frac{\ln 2}{t_{1/2}} = \frac{\ln 2}{8.04 \text{ days}} = 0.0862 \text{ days}^{-1}$$

$$\ln(N_t) = -kt + \ln(N_0)$$

$$\ln(N_t) = -(0.0862 \text{ days}^{-1})(13.8 \text{ days}) + \ln(1.25 \text{ mg}) = -0.823$$

$$N_t = e^{-0.823} = \mathbf{0.440 \text{ mg}}$$

17. Which one of the following devices converts radioactive emissions to light for detection?

- A) Geiger counter
 B) photomultiplier tube
 C) scintillation counter
 D) none of these

18. Which one of the following is true?

- A) Most spontaneous nuclear reactions are exothermic.
 B) Some spontaneous nuclear reactions are endothermic.
 C) All spontaneous nuclear reactions are exothermic.
 D) There is no relationship between exothermicity and spontaneity in nuclear reactions.

19. Binding energy per nucleon is greatest for nuclei with...

- A) large mass numbers
 B) intermediate mass numbers
 C) small mass numbers

20. When two atoms of ^{235}U are fused to form one atom of ^{238}U , the total energy involved is $3.83 \times 10^{-10} \text{ J}$. What is the total change in mass for this reaction? ($c = 3.00 \times 10^8 \text{ m/s}$)

- A) $1.28 \times 10^{-17} \text{ g}$
 B) $4.26 \times 10^{-18} \text{ g}$
 C) $1.28 \times 10^{-18} \text{ g}$
 D) $4.26 \times 10^{-19} \text{ g}$

$$E = mc^2 \quad \Rightarrow \quad m = \frac{E}{c^2} = \frac{3.83 \times 10^{-10} \text{ J}}{(3.00 \times 10^8 \text{ m/s})^2} = 4.26 \times 10^{-19} \text{ kg} = \left(\frac{1000 \text{ g}}{1 \text{ kg}}\right) = \mathbf{4.26 \times 10^{-18} \text{ g}}$$

Study Nuclear Chemistry Answer Key

J Dewey



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Table of Contents Study Nuclear Chemistry Answer Key

1. Understanding the eBook Study Nuclear Chemistry Answer Key
 - The Rise of Digital Reading Study Nuclear Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Nuclear Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Nuclear Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Nuclear Chemistry Answer Key
 - Personalized Recommendations
 - Study Nuclear Chemistry Answer Key User Reviews and Ratings
 - Study Nuclear Chemistry Answer Key and Bestseller Lists

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 - Study Nuclear Chemistry Answer Key Enhanced eBook Features
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 - Highlighting and Note-Taking Study Nuclear Chemistry Answer Key
 - Interactive Elements Study Nuclear Chemistry Answer Key
8. Staying Engaged with Study Nuclear Chemistry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Nuclear Chemistry Answer Key
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 - Creating a Diverse Reading Collection Study Nuclear Chemistry Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Nuclear Chemistry Answer Key
 - Setting Reading Goals Study Nuclear Chemistry Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Nuclear Chemistry Answer Key
 - Fact-Checking eBook Content of Study Nuclear Chemistry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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