

18. ^{137}I has a half-life of 8.04 days. Assuming you start with a 1.25 mg sample of ^{137}I , how much will remain after 13.8 days?

- A) 0.625 mg
 B) 0.268 mg
 C) 0.422 mg
 D) 0.300 mg

$$k = \frac{\ln 2}{t_{1/2}} = \frac{\ln 2}{8.04 \text{ days}} = 0.0862 \text{ days}^{-1}$$

$$\ln(N_t) = -kt + \ln(N_0)$$

$$\ln(N_t) = -(0.0862 \text{ days}^{-1})(13.8 \text{ days}) + \ln(1.25 \text{ mg}) = -0.823$$

$$N_t = e^{-0.823} = \mathbf{0.440 \text{ mg}}$$

19. Which one of the following devices converts radioactive emissions to light for detection?

- A) Geiger counter
 B) photomultiplier tube
 C) scintillation counter
 D) none of these

20. Which one of the following is true?

- A) Most spontaneous nuclear reactions are exothermic.
 B) Some spontaneous nuclear reactions are endothermic.
 C) All spontaneous nuclear reactions are exothermic.
 D) There is no relationship between exothermicity and spontaneity in nuclear reactions.

21. Binding energy per nucleon is greatest for nuclei with...

- A) large mass numbers
 B) intermediate mass numbers
 C) small mass numbers

22. When two atoms of ^3H are fused to form one atom of ^4He , the total energy evolved is $3.82 \times 10^{-12} \text{ J}$. What is the total change in mass for this reaction? ($c = 3.00 \times 10^8 \text{ m/s}$)

- A) $1.28 \times 10^{-17} \text{ g}$
 B) $4.26 \times 10^{-17} \text{ g}$
 C) $1.28 \times 10^{-16} \text{ g}$
 D) $4.26 \times 10^{-16} \text{ g}$

$$E = mc^2 \quad \Rightarrow \quad m = \frac{E}{c^2} = \frac{3.82 \times 10^{-12} \text{ J}}{(3.00 \times 10^8 \text{ m/s})^2} = 4.26 \times 10^{-17} \text{ kg} = \left(\frac{1000 \text{ g}}{1 \text{ kg}}\right) = \mathbf{4.26 \times 10^{-16} \text{ g}}$$

Study Nuclear Chemistry Answer Key

J Dewey



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