

# Campbell Biology Concepts and Connections Chapter 7 - Study Questions

One of the important waste products of photosynthesis is oxygen gas. Where does this oxygen gas come from? - correct answer Oxygen is released when water is broken down to extract electrons and protons for use in the light reactions.

The light reactions produce \_\_\_\_\_. - correct answer oxygen, NADPH, and ATP

The thylakoid membranes bring together the components necessary to carry out the \_\_\_\_\_. - correct answer light reactions

Incoming photons of light energy initiate photosynthesis by \_\_\_\_\_. - correct answer exciting electrons in pigment molecules within the photosystems, raising them to a higher energy level

What summarizes the inputs and outputs of the Calvin cycle? - correct answer  $ATP + NADPH + 3CO_2 \rightarrow G3P$

What is the equation that represents photosynthesis? - correct answer  $6CO_2 + 6H_2O \rightarrow C_6H_{12}O_6 + 6O_2$

Where does photosynthesis take place? - correct answer in the chloroplast

What connects the two photosystems in the light reactions? - correct answer An electron transport chain

What two molecules are produced by the light reactions and used to power the Calvin cycle? - correct answer ATP and NADPH

What provides electrons for the light reactions? - correct answer  $H_2O$

What provides the carbon atoms that are incorporated into sugar molecules in the Calvin cycle? - correct answer Carbon dioxide ( $CO_2$ )

What transports electrons from the light reactions to the Calvin cycle? - correct answer NADPH

The light reactions take place in the \_\_\_\_\_ and the Calvin cycle takes place in the \_\_\_\_\_. - correct answer thylakoids; stroma

# Biology Concepts And Connections Study

**M Lipman**



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reaction a redox reaction explain your answer  $2\text{NaCl (aq)} + \text{Pb (s)} \rightarrow 2\text{NaNO}_3\text{ (aq)} + \text{PbCl}_2\text{ (s)}$

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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a Al b  $\text{Ba}^{2+}$  c  $\text{Br}_2$  d Ca e  $\text{Ga}^{3+}$  f H g H

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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 what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is  
 a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe}^{2+} + \text{CO} \rightarrow \text{CO}_2 + \text{Fe}$   
 oxidation  $\text{CO} \rightarrow \text{CO}_2 + 2\text{e}^-$  reduction  $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni} \rightarrow \text{Ni}_3 + 3\text{Ag}$  oxidation reduction 25  $\text{Cu} + \text{Pb} \rightarrow \text{Pb}_2 + \text{Cu}$  oxidation reduction 26  $\text{O}_2$

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 magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a  $\text{I}_2 + \text{O}_5$   
 $5\text{CO} + \text{I}_2 \rightarrow 5\text{CO}_2 + \text{I}_2$   $\text{O}_5 + 5\text{C} + \text{O}_2 \rightarrow 5\text{CO}_2$

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 state and oxidation number concepts to identify redox reactions and to keep track of electrons

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 reducing agent also  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr}^0$  to  $\text{Sr}^{2+}$  oxidized reducing agent  $\text{O}_2$  to  $\text{O}^{2-}$  reduced ox ag  $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li}^0$  to  $\text{Li}^{+1}$  oxidized red ag  
 $\text{S}^0$  to  $\text{S}^{2-}$  reduced ox ag  $3\text{C} + \text{Br}_2$