

UNIT III - THERMOCHEMISTRY

1. Calculate the H-F bond energy given the following data.



$$\Delta H_{\text{rxn}} = -544 \text{ kJ}$$

- A. 1130 kJ/mol B. 565 kJ/mol **C. 283 J/mol** D. 300 kJ/mol E. 134 kJ/mol
2. 1.0 g of the rocket fuel hydrazine (N_2H_4) is burned in a bomb calorimeter. The temperature of the bomb calorimeter changes from 24.63 °C to 28.16 °C. The heat capacity of the bomb calorimeter is 837 J/K. Calculate the heat of combustion per mol of hydrazine in the bomb calorimeter.
- A. 95.0 kJ/mol** B. 2.96 kJ/mol C. 3.55×10^3
D. 92.4 J/mol E. 1.13×10^5 kJ/mol
3. Calculate the enthalpy change for the following reaction of nitric oxide NO with ozone O_3 .
- $$\text{NO(g)} + \text{O}_3\text{(g)} \rightarrow \text{NO}_2\text{(g)} + \text{O}_2\text{(g)}$$
- Given the following data:
- | | |
|---|--------------------------------|
| $\text{NO (g)} + \text{O (g)} \rightarrow \text{NO}_2 \text{(g)}$ | $\Delta H = -304.1 \text{ kJ}$ |
| $2\text{O}_3 \text{(g)} \rightarrow 3\text{O}_2 \text{(g)}$ | $\Delta H = -284.6 \text{ kJ}$ |
| $2\text{O (g)} \rightarrow \text{O}_2 \text{(g)}$ | $\Delta H = -495.0 \text{ kJ}$ |
- A. -1083.7 kJ B. -693.8 kJ C. -198.8 kJ **D. -93.7 kJ** E. -93.7 kJ
4. An endothermic reaction does no work on the surroundings. What values do you predict for the heat and energy change of this reaction?
- A. $q < 0, \Delta E > 0$ **B. $q < 0, \Delta E < 0$** C. $q > 0, \Delta E > 0$
D. $q > 0, \Delta E < 0$ E. $q > 0, \Delta E = 0$
5. M&M candies contain 209.0 Calories per 1.5 oz serving. The fat content accounts for 38.3% of these Calories. What is the energy of the fat content in kJ?
- A. 19.1 kJ B. 874 kJ **C. 1.91×10^{-2} kJ**
D. 3.35×10^5 kJ E. 335 kJ
6. Calculate the heat released when 2.00 g of butane from a cigarette lighter is burned.
- $$2 \text{C}_4\text{H}_{10} \text{(l)} + 13 \text{O}_2 \text{(g)} \rightarrow 8 \text{CO}_2 \text{(g)} + 10 \text{H}_2\text{O (l)}$$
- A. -90.7 kJ **B. -98.3 kJ** C. -49.2 kJ D. -108.4 kJ E. -196.6 kJ
7. When Karl Kaveman adds chilled grog to a granite mug, 10.9 kJ of energy is transferred from the mug to the grog. If the mug has a mass of 625 g and an initial temperature of 25 °C, what is the final temperature of the mug? The specific heat of granite is 0.79 J/g·K.
- A. 2.9 °C B. 14 °C **C. 22 °C** D. 25 °C E. 47 °C

General Chemistry Questions Answers

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Table of Contents General Chemistry Questions Answers

1. Understanding the eBook General Chemistry Questions Answers
 - The Rise of Digital Reading General Chemistry Questions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Questions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Questions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Questions Answers
 - Personalized Recommendations

- General Chemistry Questions Answers User Reviews and Ratings
- General Chemistry Questions Answers and Bestseller Lists
- 5. Accessing General Chemistry Questions Answers Free and Paid eBooks
 - General Chemistry Questions Answers Public Domain eBooks
 - General Chemistry Questions Answers eBook Subscription Services
 - General Chemistry Questions Answers Budget-Friendly Options
- 6. Navigating General Chemistry Questions Answers eBook Formats
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 - General Chemistry Questions Answers Compatibility with Devices
 - General Chemistry Questions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Questions Answers
 - Highlighting and Note-Taking General Chemistry Questions Answers
 - Interactive Elements General Chemistry Questions Answers
- 8. Staying Engaged with General Chemistry Questions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Questions Answers
- 9. Balancing eBooks and Physical Books General Chemistry Questions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Questions Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Questions Answers
 - Setting Reading Goals General Chemistry Questions Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Questions Answers
 - Fact-Checking eBook Content of General Chemistry Questions Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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