

Thermodynamics I

Final exam

Thermal physics

QUESTION NO. 1 (20 points)

- What is meant by an ideal gas?
 - Carbon dioxide gas is contained within a piston cylinder device initially at 20°C and $0.24\text{m}^3/\text{kg}$. It undergoes an isothermal process to state 2, where the specific volume is $0.12\text{ m}^3/\text{kg}$. It is then expanded at constant pressure to state 3, where the specific volume is $0.36\text{m}^3/\text{kg}$. Determine the initial pressure and the final pressure and temperature. Calculate the amount of work done and heat transfer through each process, in J/kg .
-

QUESTION NO. 2 (20points)

- A gas mixture with the partial pressure of its components given as follows:

Mixture component	Partial pressure kPa
	100
O ₂	80
H ₂ O	20

Determine for this mixture:

- The apparent molecular weight (M_m).
 - The gas constant (R_m).
 - The specific heat at constant pressure (c_{pm}).
 - The specific heat at constant volume (c_{vm}).
 - The specific heat ratio (γ_m).
- Propanol ($\text{C}_3\text{H}_7\text{OH}$) is burned with 50 percent excess air. Write the balanced reaction equation for complete combustion and determine the air-to-fuel ratio, on mass basis.
-

QUESTION NO. 3 (15 points)

- Explain the components of the steam power plant and apply the first law of thermodynamics on the whole cycle as well as on every component.
- Steam flows steadily through an adiabatic turbine. The inlet conditions of the steam are 10 MPa , 450°C , and 80 m/s and the exit conditions are 10 kPa , 92 percent quality, and 5m/s . The mass flow rate of the steam is 12 kg/s . Determine, the power output and the turbine inlet and exit cross sectional area.

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Lei Huang



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