

Calorimetry Lab (aka: "Burn Lab")

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Introduction:

By measuring the heat that a chemical energy source or molecular bond is released as heat and light. The heat can be measured in some indirect manner (Thermocouples used that sense temperature heat, but with caution the heat industry). The more Calories a food contains, the more heat is produced when burned. Foods high in carbohydrates will release large amounts of energy. One gram of a particular nutrient for heat release that compares to that

is called as the amount of heat (energy) required to increase the temperature of 1 gram of water by 1 °C. This process is the basis of the technique of calorimetry. In thermal investigation, you will use the methods of calorimetry to approximate the amount of energy contained in a particular food unit or other food item. This heat goes off from the calorimeter as indicated by water that is evaporated above the burning food item. This method overestimates the amount of heat given off by the food.

Purpose:

1. To determine how many calories of energy burning food items contain & compare to real number as indicated by the manufacturers?
2. Determine how this experiment can be improved to produce more accurate results.

Materials:

- | | | | |
|------------------------|------------|-------------------------|------------------|
| a) Test tubes & holder | b) Baggies | c) Thermometer or scale | d) matches |
| e) Copper-impaled | f) matches | g) Food items | h) 1 Paper clips |

Procedure:

1. Suspend your food item above the evaporating dish using 1 paper clip. The paper impaled (stratified) horizontal "vertical" method will be used for your food.
2. Measure the mass of the test tube with the copper-impaled paper clip set up above the evaporating dish as described.
3. Weigh 50.0mL of distilled water and pour it into the test tube. Record the mass of H₂O in the test tube (temperature = 1 mL of H₂O = 1 g).
4. Place the clip in the evaporating dish on the base of the ring stand. Light the test tube. Light as fast as it will be lit, then place the clip. Check the temperature of the water and record.
5. Measure accurately the initial temperature of the water.
6. After 5 minutes, move the impaled test tube under the test tube. Use a match to ignite the bottom of the test tube and then put the burning test tube under the test tube. The idea is to get as much of the heat from the burning test tube as you can get from the water in the test tube.
7. Light the test tube with the match in the test tube under the test tube. Measure the highest temperature of the water and record it in the data table. No, light it if it goes out.
8. Measure the final mass of the test tube in the evaporating dish of the burning (before data).
9. Repeat steps 1-7 using a different food item.

Calorimetry Lab Aka Burn Answers

Wolfgang Guggemos



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