

Generalmente si esprime come il rapporto  $\frac{r_{anione}}{r_{catione}}$ , dove  $r$  indica il raggio dello ione. Questo rapporto dipende dagli ioni coinvolti negli ioni degli ioni.

Numero di coordinazione	Geometria	Intervallo di rapporti $r_{anione}/r_{catione}$	Struttura cristallina
2	Lineare	0,15 - 0,22	Struttura di coordinazione
4	Tetraedrica	0,22 - 0,41	Struttura di coordinazione
6	Octaedrica	0,41 - 0,73	Struttura di coordinazione
8	Cubica	0,73 - 1,00	Struttura di coordinazione

Il rapporto  $r_{anione}/r_{catione}$  determina il numero di coordinazione.

**nelle molecole:**

- 1) Sono molecolari, possono legarsi tra loro in modo a legami a H.
- 2) Sono con struttura cristallina molecolare (atomi legati tra legami covalenti) cristalli molecolari,  $SiO_2$ ,  $SiH_4$ ,  $Si_2H_6$ .
- 3) La struttura cristallina molecolare è tridimensionale e non è cristallina, ma può essere cristallina se legati a una sola direzione. Se la struttura è bidimensionale, si cristallizza in strati. Se tridimensionale, si cristallizza in strati o in modo tridimensionale.
- 4) La struttura cristallina molecolare è tridimensionale e non è cristallina, ma può essere cristallina se legati a una sola direzione.
- 5) La struttura cristallina molecolare è tridimensionale e non è cristallina, ma può essere cristallina se legati a una sola direzione.



Struttura del ghiaccio. Non considerando gli H legati all'ossigeno, si è ottenuto un reticolo di ossigeno di un cristallo di un cristallo. Questo significa che il cristallo di ghiaccio è un cristallo di ossigeno.

La struttura cristallina molecolare è tridimensionale e non è cristallina, ma può essere cristallina se legati a una sola direzione.

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Il numero di coordinazione è 4 per il reticolo cubico e 3 per il reticolo esagonale.

- 1) Sono molecolari, possono legarsi tra loro in modo a legami a H.
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**proprietà dei solidi**

- 1) Sono molecolari, possono legarsi tra loro in modo a legami a H.
- 2) Sono con struttura cristallina molecolare (atomi legati tra legami covalenti) cristalli molecolari,  $SiO_2$ ,  $SiH_4$ ,  $Si_2H_6$ .
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**SJ Ball**



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