

CHAPTER 2

SOLID-STATE CHEMISTRY

Of the three states of matter, solids possess the most structural diversity. Whereas gases and liquids consist of discrete molecules that are randomly distributed due to thermal motion, solids consist of molecules, atoms, or ions that are statically positioned. To fully understand the properties of solid materials, one must have a thorough knowledge of the structural interactions between the subunit atoms, ions, and molecules. This chapter will outline the various types of solids, including structural classifications and nomenclature for both crystalline and amorphous solids. The material in this key chapter will set the groundwork for the rest of this textbook, which describes a variety of materials classes.

2.1. AMORPHOUS VS. CRYSTALLINE SOLIDS

A solid is a material that retains both its shape and volume over time. If a solid possesses long range, regularly repeating units, it is classified as a *crystalline* material. Crystalline solids are only produced when the atoms, ions, or molecules have an opportunity to organize themselves into regular arrangements, or *lattices*. For example, crystalline minerals found in nature have been formed through many years of extreme temperature and pressure, or slow evaporation processes. Most naturally occurring crystalline solids comprise an agglomeration of individual microcrystalline units; single crystals without significant defects are extremely rare in nature, and require special growth techniques (see p. 28).

If there is no long-range structural order throughout the solid, the material is best described as *amorphous*. Quite often, these materials possess considerable short-range order over distances of 1–10 nm or so. However, the lack of long-range translational order (periodicity) separates this class of materials from their crystalline counterparts. Since the majority of studies have been addressed to study crystalline solids relative to their amorphous counterparts, there is a common misconception that most solids are crystalline in nature. In fact, a solid product generated from many chemical reactions will be amorphous by default, unless special procedures are used to facilitate molecular ordering (*i.e.*, crystal formation). Although the crystalline state is more thermodynamically-favorable than the

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