

## Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

### Chemical Bonding:

- **Four types of bonds:**
  - *Nonpolar Covalent Bond* ( $\Delta EN$  (difference in electronegativity) = < 0.3)
    - **EQUAL SHARING** of electrons between the 2 bonding nonmetal atoms
  - *Polar Covalent Bond* ( $\Delta EN = 0.4-1.7$ )
    - **UNEQUAL SHARING** of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
  - *Ionic Bond* ( $\Delta EN = > 1.7$ )
    - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely **GIVING UP/TRANSFERRING ELECTRONS** to the atom with the greater EN value
    - Ionic bond results due to the attraction between the **POSITIVE** cation (formed by losing electrons) and the **NEGATIVE** anion (formed by gaining electrons)
  - *Metallic Bond*
    - Bonding between two metallic atoms which occurs due to the attraction between a “sea of electrons” and the nuclei of bonded metal atoms
- \* Must memorize/know the difference in electronegativity ranges for each type of bond
  - \*\* Note: you should still know the definition of electronegativity (EN)
- **Properties of the Four Types of Chemical Bonds**
  - See the table at the bottom of Packet pg.5
  - You **MUST** be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness
- **Why & How do chemical bonds form?**
  - **Why** → because all atoms want to be at their **LOWEST POTENTIAL ENERGY**; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
  - **How** → **ATTRACTIVE** and **REPULSIVE** forces
    - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
    - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repel one another (because they have the same charge)
  - **Bond Length**
    - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
  - **Bond Energy**
    - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

# Chapter 6 Study Guide Chemistry

**ML Yell**



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