



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - C_2H_6
 - H_2SO_4
 - CaCl_2
 - $\text{C}_2\text{H}_5\text{OH}$
 - H_2O_2
 - C_2H_4
 - C_2H_2
 - C_2H_6

- Classify each of the following as either ionic or molecular compound. Write the name of the compound.
- Write the formula for each of the following molecular compounds. Write the name of the compound.
- Write the name of each of the following ionic compounds. Write the formula of each compound.

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride. $\text{H}-\text{F}$
- Draw the Lewis structure for carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Draw the Lewis structure for methane. C
- Draw the Lewis structure for ethane. $\text{H}-\text{C}-\text{C}-\text{H}$
- Draw the Lewis structure for acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for ammonia. N



SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital do you combine to form the sigma bond in the hydrogen molecule and ethane?
- Which orbital and sub-orbital give rise to the pi bond in ethene and ethyne?

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Chemistry John S. Phillips, Cheryl Wistrom, 2000 **Science Citation Index** ,1992 Vols for 1964 have guides and journal lists

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