

Chapter Test B

Chapter: Acid-Base Titration and pH

PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The pH scale generally ranges from
- 0 to 1.
 - 1 to 1.
 - 0 to 7.
 - 0 to 14.
- _____ 2. During the titration of HCl with NaOH, a very rapid change in pH occurs
- when the first addition of known solution is made.
 - when roughly equivalent amounts of H_3O^+ ions and OH^- ions become present.
 - at several points.
 - at no point.
- _____ 3. A water solution is neutral if
- it contains no H_3O^+ ions.
 - it contains no ionized water molecules.
 - it contains no H_3O^+ ions or OH^- ions.
 - the concentrations of H_3O^+ ions and OH^- ions are equal.
- _____ 4. The antilog of a number N is
- the inverse of N .
 - the square root of N .
 - 10 raised to the power of N .
 - N raised to the 10th power.
- _____ 5. Universal indicators
- are mixtures of several indicator solutions.
 - are pure substances.
 - have very brief color-change intervals.
 - work well only for acidic solutions.
- _____ 6. A useful pH range for an indicator in neutralizations involving strong acids and weak bases is
- 1.2 to 3.0.
 - 3.1 to 4.6.
 - 6.0 to 7.6.
 - 9.5 to 11.0.

Chapter Test B Chemistry Answers

Alexander Smith



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Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science ,2003-11
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Dental Aptitude Test Emanuel C. Gruber,Edward C. Gruber,1963

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