

NUCLEAR CHEMISTRY

Practice Problems

In your notebook, solve the following problems.

SECTION 25.1 NUCLEAR RADIATION

- What happens to the mass number and atomic number of an atom that undergoes beta decay?
- A nuclide of an element undergoes alpha particle decay. How do the atomic number and mass number of the particle change?
- Give the composition of the nucleus of the following isotopes.
 - ${}^{234}\text{Th}$
 - ${}^{232}\text{Th}$
 - ${}^{210}\text{Po}$
- Complete each of the following equations.
 - ${}^{234}\text{Th} \rightarrow \text{?} + \text{?}$
 - ${}^{226}\text{Ra} \rightarrow \text{?} + \text{?}$
 - ${}^{234}\text{Th} \rightarrow \text{?} + \text{?}$

SECTION 25.2 NUCLEAR TRANSFORMATIONS

- Write a nuclear equation for the following radioactive processes.
 - alpha decay of francium-223
 - electron capture by beryllium-7
 - beta emission by argon-37
 - positron emission by boron-12
- Complete the equations for these transmutation reactions.
 - ${}^{235}\text{U} + \text{?} \rightarrow \text{?} + \text{?}$
 - ${}^{238}\text{U} + \text{?} \rightarrow \text{?} + \text{?} + \text{?}$
 - ${}^{238}\text{U} + \text{?} \rightarrow \text{?} + \text{?} + \text{?}$
 - ${}^{238}\text{U} \rightarrow \text{?} + \text{?} + \text{?} + \text{?}$
 - $\text{?} + \text{?} \rightarrow \text{?} + \text{?} + \text{?} + \text{?}$
- Polonium-214 has a relatively short half-life of 264 s. How many seconds would it take for 6.0 g of this isotope to decay to 0.25 g?
- How many days does it take for 16 g of gadolinium-153 to decay to 1.0 g? The half-life of gadolinium-153 is 17 days.
- By approximately what factor would the mass of a sample of copper-66 decrease in 10 minutes? The half-life of copper-66 is 5.49 min.
- In 5.49 seconds, 1.00 g of argon-39 decay to leave only 0.10 g. What is the half-life of argon-39?

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David Baud



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