

NUCLEAR CHEMISTRY WORKSHEET

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Most nuclear reactions involve changes in the number of protons and/or neutrons in the nucleus of an atom. These changes are called *transmutations* because an atom of one element is changed into an atom of a different element. During these changes, nuclear particles may be absorbed and/or emitted along with a release of energy.

There are four types of nuclear reactions:

- 1) *natural radioactive decay*: transmutations occur by natural radioactivity (the ability of a nucleus to emit a nuclear particle and energy without external stimulation);
- 2) *artificial transmutation*: transmutations occur by nuclear disintegration caused by external stimulation as a scientist bombards a nucleus with a particle (the addition of another nuclear particle makes the nucleus unstable);
- 3) *fission*: certain nuclei having a large mass are bombarded with special particles that cause the nuclei to split into two nuclei each having a smaller mass;
- 4) *fusion*: nuclei of light elements are combined to form heavier nuclei.

Nuclear reactions are written using symbols in the A_ZX notation. The symbols for the major nuclear particles involved in nuclear reactions are given below.

Alpha particle	${}^4_2\text{He}$ (an alpha particle is a helium nucleus)
Proton	${}^1_1\text{H}$ (the most common hydrogen nucleus is a proton)
Neutron	${}^1_0\text{n}$
Electron	${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$ (also called a beta particle)
Positron	${}^0_1\text{e}$ or ${}^0_1\beta$
Gamma ray	${}^0_0\gamma$

The table below contains information about nuclear reactions involving the emission of radiation.

NUCLEAR CHANGES

Type of Reaction	Radiation (Particle)	Effect on the Atomic Number	Effect on the Atomic Mass
Alpha emission (α)	${}^4_2\text{He}$	decrease by 2	decrease by 4
Beta emission (β)	${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$	increase by 1	no change
Positron emission (β^+)	${}^0_1\text{e}$ or ${}^0_1\beta$	decrease by 1	no change
Gamma emission	${}^0_0\gamma$	no change	no change
Electron capture (EC)	${}^0_{-1}\text{e}$ or ${}^0_{-1}\beta$	decrease by 1	no change

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**California. University. Lawrence
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Table of Contents Chapter 28 Nuclear Chemistry Worksheet Answers

1. Understanding the eBook Chapter 28 Nuclear Chemistry Worksheet Answers
 - The Rise of Digital Reading Chapter 28 Nuclear Chemistry Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 28 Nuclear Chemistry Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 28 Nuclear Chemistry Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 28 Nuclear Chemistry Worksheet Answers
 - Personalized Recommendations
 - Chapter 28 Nuclear Chemistry Worksheet Answers User Reviews and Ratings
 - Chapter 28 Nuclear Chemistry Worksheet Answers and Bestseller Lists

5. Accessing Chapter 28 Nuclear Chemistry Worksheet Answers Free and Paid eBooks
 - Chapter 28 Nuclear Chemistry Worksheet Answers Public Domain eBooks
 - Chapter 28 Nuclear Chemistry Worksheet Answers eBook Subscription Services
 - Chapter 28 Nuclear Chemistry Worksheet Answers Budget-Friendly Options
6. Navigating Chapter 28 Nuclear Chemistry Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 28 Nuclear Chemistry Worksheet Answers Compatibility with Devices
 - Chapter 28 Nuclear Chemistry Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 28 Nuclear Chemistry Worksheet Answers
 - Highlighting and Note-Taking Chapter 28 Nuclear Chemistry Worksheet Answers
 - Interactive Elements Chapter 28 Nuclear Chemistry Worksheet Answers
8. Staying Engaged with Chapter 28 Nuclear Chemistry Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 28 Nuclear Chemistry Worksheet Answers
9. Balancing eBooks and Physical Books Chapter 28 Nuclear Chemistry Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 28 Nuclear Chemistry Worksheet Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 28 Nuclear Chemistry Worksheet Answers
 - Setting Reading Goals Chapter 28 Nuclear Chemistry Worksheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 28 Nuclear Chemistry Worksheet Answers
 - Fact-Checking eBook Content of Chapter 28 Nuclear Chemistry Worksheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

-
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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