

## Chapter 5. Chemical potential and Gibbs distribution

### 1 Chemical potential

So far we have only considered systems in contact that are allowed to exchange "heat", i.e. systems in thermal contact with one another. In this chapter we consider systems that can also exchange particles with one another, i.e. systems that are in *diffusive* contact.

Consider 2 systems  $\mathcal{S}_1$  and  $\mathcal{S}_2$  that are in diffusive contact with one another and in thermal contact with a 3rd system, a reservoir at temperature  $\tau$ . We have shown that the Helmholtz free energy for the combined system  $\mathcal{S}_1 + \mathcal{S}_2$  will be a minimum when it is in equilibrium with the reservoir. We must therefore minimize

$$F = F_1 + F_2$$

with respect to the distribution of the particles between  $\mathcal{S}_1$  and  $\mathcal{S}_2$  to find the equilibrium state of this combined system. The total number of particles in the system is fixed, so that

$$dF = \left( \frac{\partial F_1}{\partial N_1} \right)_{\tau} dN_1 - \left( \frac{\partial F_2}{\partial N_2} \right)_{\tau} dN_2 = 0$$

in equilibrium, i.e.

$$\left( \frac{\partial F_1}{\partial N_1} \right)_{\tau} = \left( \frac{\partial F_2}{\partial N_2} \right)_{\tau}$$

The quantity

$$\mu(\tau, V, N) = \left( \frac{\partial F}{\partial N} \right)_{\tau, V}$$

is known as the *chemical potential*, so that our equilibrium condition is that

$$\mu_1 = \mu_2.$$

Inspecting the expression for  $dF$ , we see that when  $\mu_1 > \mu_2$  moving particles from  $\mathcal{S}_1$  to  $\mathcal{S}_2$  decreases  $F$ , taking the system closer to equilibrium. Thus, particles tend to flow from systems of high chemical potential to systems of lower chemical potential.  $\mu$  is the (Helmholtz) free energy "per particle" in a system.

If several chemical species are present within a system, then there is chemical potential associated with each distinct species, e.g.

$$\mu_j = \left( \frac{\partial F}{\partial N_j} \right)_{\tau, V, N_1, \dots, N_{j-1}}$$

is the chemical potential for species  $j$ .

#### 1.1 Example: the ideal gas

In chapter 3 we showed that the Helmholtz free energy of an ideal monatomic gas is

$$F = -N\tau \ln(n_Q V) + N\tau \ln N - N\tau,$$

so that

$$\mu = -\tau \ln(n_Q V) + \tau \ln N = \tau \ln \left( \frac{N}{n_Q} \right).$$

# Chapter 5 Chemical Potential And Gibbs Distribution 1

**Julia Schneider**



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web für kinder und jugendliche gibt es für die asthma therapie ein spezielles stufenschema das sich vor allem durch eine zusätzliche therapiestufe vom schema für erwachsene unterscheidet ziel jeder therapie ist es kindern mit asthma ein normales leben zu ermöglichen

*asthma beim kind was eltern wissen müssen asthma de* - Apr 07 2022

web novartis ist asthma erblich als häufiger risikofaktor für asthma gerade bei kindern gilt das verstärkte auftreten in der familie ist ein elternteil asthmatiker liegt das risiko bei rund 25 prozent sind vater und mutter betroffen erhöht es sich auf mehr als 50 prozent

asthmatherapie bei kindern und jugendlichen mit fixer ics - Jun 21 2023

web aug 20 2021 für jugendliche gilt in stufe 4 und 5 bei langzeittherapie mit ics und formoterol dass diese kombination dann auch bei bedarf zum einsatz kommen sollte für stufe 1 und 2 reicht die evidenz aus den großen erwachsenenlastigen

studien aus um auch bei mildem asthma eine kombination aus niedrig dosiertem ics mit rasch