

## IONIC BONDING

Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to obtain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.

1.  $\text{K} + \text{F}$



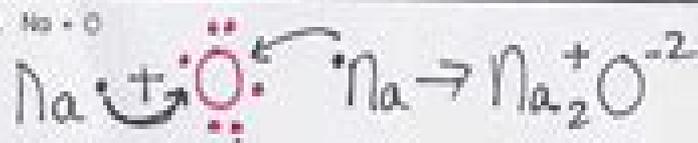
2.  $\text{Mg} + \text{I}$



3.  $\text{Be} + \text{S}$



4.  $\text{Na} + \text{O}$



5.  $\text{Al} + \text{Br}$



## COVALENT BONDING

Name \_\_\_\_\_

Covalent bonding occurs when two or more nonmetals share electrons, attempting to gain a stable octet of electrons at least part of the time. For example:

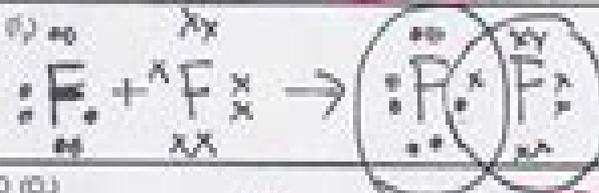


Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

1.  $\text{H} + \text{H} (\text{H}_2)$



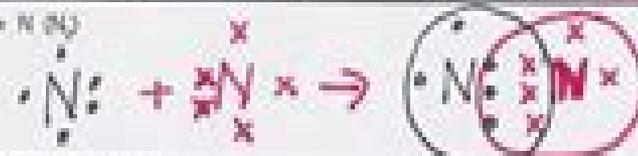
2.  $\text{F} + \text{F} (\text{F}_2)$



3.  $\text{O} + \text{O} (\text{O}_2)$



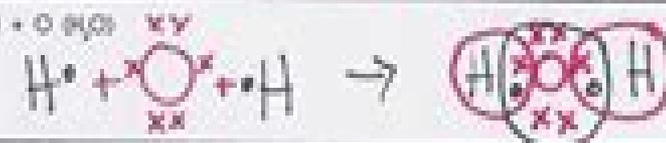
4.  $\text{N} + \text{N} (\text{N}_2)$



5.  $\text{C} + \text{O} (\text{CO}_2)$



6.  $\text{H} + \text{O} (\text{H}_2\text{O})$



# Chapter 8 Covalent Bonding Workbook Answers

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