

Section 12.1

OBJECTIVE 1

Write complete balanced redox reactions that are indicated.

1. _____
 - a. _____
 - (a) balance in terms of all the atoms and products.
 - (b) indicate oxidation-reduction reactions and products.
 - (c) number of moles of each element in each compound in reaction.
 - (d) number of electrons transferred in reaction.
 - b. _____
 - (a) indicate the following reactions are not redox reactions.
 - (b) the name of each product of each redox reaction.
 - (c) the number of moles of hydrogen that will react with a certain quantity of oxygen.
 - (d) the balanced chemical equation for the reaction.
 - c. _____
 - (a) balanced chemical equation with the state symbols.
 - (b) number of moles of each substance in the reaction.
 - (c) energy released in the reaction.
 - (d) electron configuration of all elements in the reaction.
 - (e) number of moles of each substance in the reaction.
 - d. _____
 - (a) calculate number of moles of hydrogen in each of oxygen molecules in reaction.
 - (b) number of moles of _____.
 - (c) number of moles of _____.
 - (d) number of moles of _____.
 - (e) number of moles of _____.

2. Write the balanced equation for the following redox reaction:



a. Balance the equation:



b. Write the oxidation of _____ to _____.

c. In _____ of _____ how many moles of _____ are _____?

d. Write a balanced equation of _____ to _____ as _____ of _____.

e. Write a balanced equation of _____ to _____ as _____ of _____.

Chapter 9 Review Stoichiometry Answers Section 1

G Thomas



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