

Arrangement of Electrons in Atoms

Section 2: Atomic Electronic Configurations

In this section, you will study the order of the orbitals and the rules that govern the arrangement of electrons in atoms. You will also study the Aufbau principle, Pauli exclusion principle and Hund's rule.

1. The arrangement of electrons in atoms is governed by the Aufbau principle, Pauli exclusion principle and Hund's rule. Which of the following is NOT a principle?
 - a. Hund's rule
 - b. Aufbau's rule
 - c. Pauli exclusion principle
 - d. The Pauli exclusion principle
2. According to the Aufbau principle, which of the following orbitals is filled last?
 - a. 4s
 - b. 3d
 - c. 4p
 - d. 3s
3. In a ground state hydrogen atom, which orbital is the most energetic?
 - a. 1s
 - b. 2s
 - c. 2p
 - d. 3s
4. What is the maximum number of electrons that can occupy the orbitals with the principal quantum number $n = 3$?
 - a. 9
 - b. 18
 - c. 27
 - d. 36
5. The ground state electronic configuration of iron is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$. In this configuration, what is the number of orbitals that contain only one electron?
 - a. 5
 - b. 6
 - c. 10
 - d. 11

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