

- A. $n = 4, l = 3, m_l = -2, m_s = +1/2$
- B. $n = 3, l = 2, m_l = -3, m_s = -1/2$
- C. $n = 3, l = 0, m_l = 0, m_s = +1/2$
- D. $n = 4, l = 1, m_l = 1, m_s = -1/2$
- E. $n = 2, l = 0, m_l = 0, m_s = +1/2$

39. The orbital diagram for a ground-state oxygen atom is

- A. $1s \uparrow\downarrow \quad 2s \uparrow\downarrow \quad 2p \uparrow\downarrow \uparrow _$
- B. $1s \uparrow\downarrow \quad 2s \uparrow \quad 2p \uparrow \uparrow \uparrow$
- C. $1s \uparrow\downarrow \quad 2s \uparrow\downarrow \quad 2p \uparrow \uparrow \uparrow$
- D. $1s \uparrow\downarrow \quad 2s \uparrow\downarrow \quad 2p \uparrow\downarrow \uparrow \uparrow$
- E. $1s \uparrow\downarrow \quad 2s \uparrow\downarrow \quad 2p \uparrow\downarrow \uparrow\downarrow \uparrow$

40. Which of the following ground-state electron configurations are correct for the atom or ion shown?

- A. $\text{Cl}^-: [\text{Ne}]3s^23p^4$
- B. $\text{Al}^{3+}: [\text{Ne}]$
- C. $\text{As}: [\text{Ar}]4s^24p^3$
- D. $\text{Ti}: [\text{Ar}]4s^24d^2$

41. In what row and group of the periodic table would you find the element with the electron configuration $[\text{Kr}]5s^24d^{10}5p^2$?

- A. row 4, group 5A
- B. row 4, group 4A
- C. row 5, group 5A
- D. row 5, group 4A

42. Which one of the following pairs are isoelectronic?

- A. Mn^{2+} and Ar
- B. K^+ and Cl^-
- C. Zn^{2+} and Cu^{2+}
- D. Na^+ and K^+
- E. Cl^- and S

43. Which of the atoms listed below has the largest radius?

- A. Cl
- B. I

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LL Leslie



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